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The cross-talk between lateral sheet dimensions of pristine graphene oxide nanoparticles and Ni²⁺ adsorption†

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This study investigated the removal of nickel(II) ions by using two sizes of graphene oxide nanoparticles (GO – 450 nm and GO – 200 nm). The thickness and lateral sheet dimensions of GO are considered to be an important adsorbent and promising method for sufficient removal of metals like nickel, lead, copper, etc. The graphite oxide was prepared by oxidation–reduction reaction (Hummers method), and the final product was labelled as GO – 450 nm. A tip sonicator was used to reduce the size of particles to 200 nm under controlled conditions (time and power of sonication). FTIR spectroscopy shows that both sizes of GO particles contain several types of oxygen groups distributed onto the surface of GO particles. Scanning electron microscopy (SEM) and the statistical analysis confirmed the formation of these two sizes of GO particles. The GO – 200 nm performed better removal of Ni(II) compared with GO – 450 nm, due to more surfaces being available. The adsorption capacity of GO particles increased drastically from 45 mg g^{−1} to 75 mg g^{−1} for GO – 450 nm and GO – 200 nm respectively, these values were carried out after 2 h of incubation. The kinetics of adsorption and several parameters like initial concentration at equilibrium, pH, temperature, and adsorbent dose are controlled and studied by using UV-visible spectroscopy. The results indicated a significant potential of GO – 200 nm as an adsorbent for Ni(II) ion removal. An additional experiment was performed to estimate the surface area of GO – 450 nm and GO – 200 nm, the results show that the surface areas of GO – 450 nm and GO – 200 nm are 747.8 m² g^{−1} and 1052.2 m² g^{−1} respectively.

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Introduction

Graphene oxide (GO) has been investigated as an extraordinary next-generation adsorbent in water and wastewater purification.^{1–3} The neat structure and outstanding physicochemical characteristics of GO have attracted enormous attention from researchers for use in several fields of technology, such as energy, biology, medicine, and water.^{2,4–14} GO is a two-dimensional layer of sp² hybridized carbon atoms decorated with abundant oxygen groups like carboxyl, hydroxyl, epoxy, etc., these groups have the ability to bind with organic and inorganic substances with both chemical and/or physical interactions. Moreover due to its highly hydrophilic surfaces, open flake morphology and high adsorption capacity towards different metallic ions, organic and inorganic materials, several

reports have been focused on GO as a promising nanosystem in wastewater treatment. It exhibits no obvious toxicity under low dose (*ca.* 0.2 mg) and medium dose (*ca.* 0.2–0.25 mg).^{15–17}

Heavy metals such as cadmium, cobalt, lead, nickel are non-bio-degradable and cause various diseases and disorders. Nickel is a transition metal ion that has four oxidation states (+1, +2, +3, and +4), it is widely used in industry such as mining, smelting, textiles, fertilizer, electroplating, battery manufacturing, and pigment production. At high concentrations, Ni²⁺ may lead to stern damage of vital organs such as lungs, kidneys, gastrointestinal irritation, and lung and bone cancers. The Ni²⁺ in wastewater varies from a low value of 0.5 mg l^{−1} to a high value of 1000 mg l^{−1}. The maximum permissible safe limit of Ni²⁺ with industrial effluents into land water is 3 mg l.^{3,18,19} Innovative approaches for treating industrial wastewater including heavy metals often involve new technologies for reducing poisonous in order to gather technology-based treatment standards. The incorporation of nanoparticles such as functionalized graphene and carbon nanotubes into water purification technologies against the removal of heavy metals appears as a very dynamic branch of nanotechnology. Nanoparticles owe their potential to the high specific surface area and surface reactivity compared to conventional substances. Depending on

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the mechanism of uptake, nanoparticles can be destined to establish high selectivity against various types of pollutants. Graphene oxide has a very extensive advantage in treating various heavy metal ions from aqueous solutions, such as lead (Pb(II)),^{20,21} copper (Cu(II)),^{22,23} cobalt (Co(II)),²⁴ cadmium (Cd(II)),²² chromium (Cr(VI)).²⁵ Moreover, GO functionalized by using organic molecules in order to enhance the surface selectivity of pollutants.^{22,26,27}

Nowadays, an extensive research on the facile synthesis of 2D carbon nanomaterials with inorganic substances derived from MOF precursors for various applications.^{28–30} The determination of heavy metal ions in water can be achieved using various methods. The simple and facile method is a chemical reagent method such as (4-(2-pyridylazo) resorcinol) which depends on changes in color when reacting with heavy metal ions. The more precise detection method is carried out using instruments like inductively coupled plasma-mass spectrometry (ICP-MS), atomic absorption spectrophotometer (AAS), and X-ray fluorescence spectrometry (XRF). However, those techniques require additional chemicals, expensive equipment, and trained users. Thus, in this article, ultraviolet/visible (UV/VIS) spectroscopy method was employed as a non-destructive method for the detection of nickel(II) without chemical reagents.³¹

To the best of our knowledge, no studies have been conducted to investigate the effect of lateral sheet dimensions of pristine GO on the adsorption of Ni²⁺ in solution, as well as no studies joint between the size distribution and surface area of GO nanoparticles. We have investigated two sizes of GO nanoparticles (GO – 450 nm and GO – 200 nm) at different time intervals, concentration, pH value, adsorbent dose and temperature.

Materials and methods

Synthesis of graphite oxide

As shown in Scheme 1 (ESI[†]), graphite oxide was prepared by using the Hummer's method and conventionally by the oxidation–exfoliation reaction.^{32–34} Briefly, 1.0 g of graphite and 50 g NaCl were grounded in mortar for 20 min, the ground graphite was dissolved in distilled water, filtrated, washed several times and dried in an oven at 40 °C for 6 h. The filtrated graphite was stirred in 23 ml 95% H₂SO₄ overnight. The mixture was placed in an ice bath (below 10 °C) for starting the oxidation step, 3.0 g

KMnO₄ was added slowly over 3 hours with continuous stirring. Afterward, the mixture was heated up to 35 °C for 30 min and to 50 °C for 45 min respectively. The mixture cooled down then 46 ml of distilled water was slowly added to the solution, the solution temperature was increased to 98–105 °C with stirring for 45 min. A 140 ml distilled water and 10 ml of 32% H₂O₂ were added gradually to terminate the reaction. The final product was filtrated and washed 5 times with 5% HCl and distilled water. Lastly, the graphite oxide was dried in an oven at (50 °C) for 6 h.³⁴ The actual yield of GO is 1.46 g.

Synthesis of graphene oxide nanoparticles

The graphene oxide nanoparticles were prepared according to our previous protocol.³⁴ Shortly, 1.0 mg ml^{−1} of graphite oxide was sonicated in an ultra-sonication bath under controlled conditions (power, concentration and time), as shown in Scheme 1,† step no. 3 (Size reduction).

Preparation of different nickel(II) concentrations

A 1000 mg l^{−1} stock solution of NiSO₄·6H₂O was prepared by dissolving 0.4478 g in 100 ml distilled water. The stock solution was used for the preparation of diluted solutions (200, 300, 400, 500, 600, 700, 800, and 900 ppm).

Batch adsorption experiments

Batch adsorption experiments were carried out in a water bath sonicator. The effect of adsorbent dose, contact time, initial concentration, pH, and temperature were studied by using a UV-Vis spectrophotometer. Two sizes of GO particles (450 nm and 200 nm) were used as an adsorbent to understand the cross-talk between different sizes of GO and Ni²⁺ adsorption. A 3.0 ml nickel solution of different concentrations was shaken in closed bottles with 20 mg GO as the desired adsorbent dose for various contact time (10–120 min) and pH values from 2 to 10. The separation of the solid phase from a solution was done by using suction filtration, and then the residual nickel amount in the filtrate was determined.

The surface area of GO nanoparticles was determined by using methylene blue (MB) experiment.³⁵ A 100 ppm of methylene blue dye as stock solution was prepared in ultrahigh distilled water, different solutions of methylene blue were

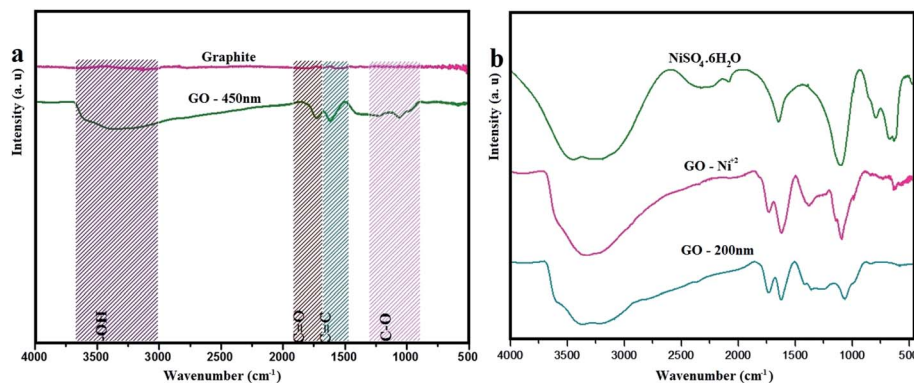


Fig. 1 FTIR spectra of (a) graphite and GO – 450 nm. (b) NiSO₄·6H₂O, GO – Ni²⁺ and GO – 200 nm.

prepared for establishing the calibration curve. In separate experiments, 5 mg of GO was mixed with 25 ml of 16, 8, 4, 2 and 1 ppm of MB for 10 min at 20 °C and 4000 rpm. The samples were filtrated and the concentration of MB at equilibrium was calculated.

Result and discussion

Scanning electron microscopy (SEM)

The scanning electron microscopy (SEM) is employed to study the morphology and the lateral sizes of GO nanoparticles.^{34,36} As shown in Fig. 2, SEM exhibits two sizes of graphene oxide nanoparticles; as shown in panel (a) the average size (lateral width) of as-prepared GO particles is approximately 450 nm. The average size of GO particles after sonication by using tip sonicator is approximately 200 nm as shown in panel (b). Fig. 2c represents the statistical analysis of GO particles deduced from SEM images. The SEM images were obtained using a FEI, NOVA NanoSEM-200 with an acceleration voltage of 15 kV.

FTIR

Fig. 1 exhibits the FTIR spectra of graphite and as prepared graphite oxide (GO – 450 nm). The graphite spectrum shows no notable bands in the region of IR from 4000–500 cm^{-1} . However GO reveals variation bands at 3312, 1730, 1612, 1231, 1077 cm^{-1} which corresponding to hydroxyl, carbonyl, $\text{C}=\text{C}$, epoxy, and $\text{C}-\text{O}$ groups respectively.² The broad band at 3312 cm^{-1}

attributed to the stretching vibration of ($-\text{OH}$) group, which is due to the hydroxyl group of carboxylic acid at the edges of GO sheets as well as the alcohol groups distributed on the basal plane of graphene oxide layers. The significant peak at 1612 cm^{-1} ascribed to sp^2 aromatic pattern ($\text{C}=\text{C}$), and the peak appeared at 1730 cm^{-1} is due to ($\text{C}=\text{O}$) group. The weak band in the region of 1231 cm^{-1} is assigned to stretching vibration of epoxide group ($\text{C}-\text{O}$), and the peak at 1077 cm^{-1} due to the alkoxy group ($\text{C}-\text{O}$).^{24,25} Fig. 1b shows the FTIR spectra of free Ni^{2+} , GO – Ni^{2+} , and GO – 200 nm. Ni^{2+} reveals a broad peak at 3200 cm^{-1} , which assigned to symmetric stretching of the water molecule, moreover the bending vibration of water molecules observed at 1656 cm^{-1} . The other peaks at 1096 cm^{-1} , 465 cm^{-1} , 632 cm^{-1} , and 797 cm^{-1} are corresponding to fundamental vibrations of SO_4^{2-} free ion namely as a non-degenerate, doubly degenerate, triply degenerate mode respectively. After adsorption of Ni^{2+} , a broad band at 3312 cm^{-1} in GO was shifted to 3150 cm^{-1} . The strong peak at 1730 cm^{-1} shifted to 1715 cm^{-1} , which is assigned to Ni^{2+} adsorbed on GO particles. These findings are in agreement with previously reported data.^{37,38} The FTIR spectra were obtained using Bruker FTIR in the range between 4000–500 cm^{-1} .

UV-visible spectrophotometry

The UV-visible Spectrophotometry (UV-Visible line 9100, photometric range: 320–1100 nm, Aqualabo company) was used in this study. Fig. 1a (ESI[†]) reveals the UV visible spectroscopy of

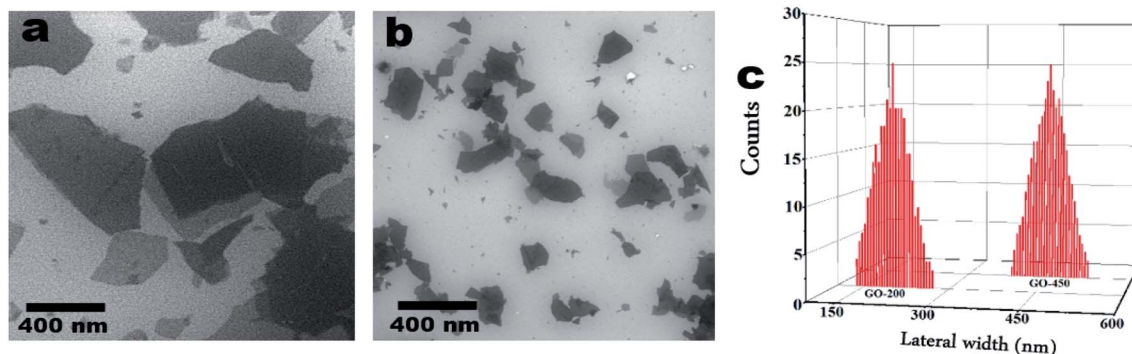


Fig. 2 SEM images of (a) GO – 450 nm and (b) 200 nm. Panel (c) is the average width (nm) of GO particles deduced from SEM images.

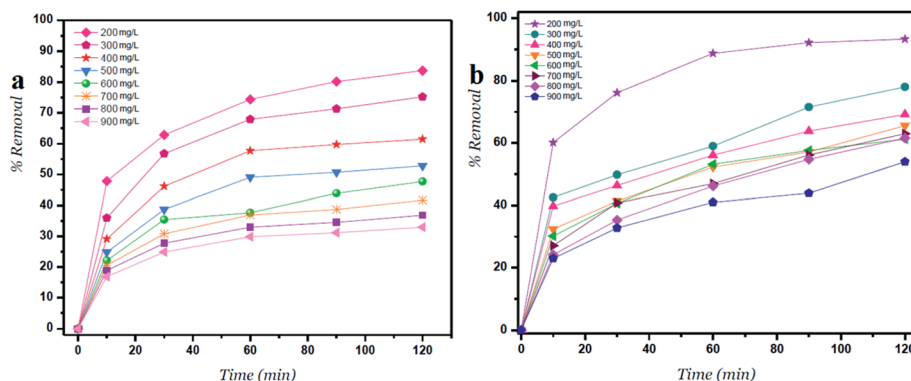


Fig. 3 Percent removal versus time for (a) GO – 450 nm and (b) GO – 200 nm at different concentrations.



free Ni^{2+} ions at different concentrations. As the Ni^{2+} concentration increases the absorbance increases at different wavelengths (393 nm, 656 nm, and 720 nm). These wavelengths are assigned to $\nu_1 = {}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^1\text{E}_\text{g}$, $\nu_2 = {}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^3\text{T}_1\text{g}(\text{F})$, $\nu_3 = {}^3\text{A}_2\text{g}(\text{F}) \rightarrow {}^3\text{T}_1\text{g}(\text{P})$. The observed bands of 393 nm and 720 nm related to ν_3 and ν_2 respectively. While ν_1 refers to 656 nm.³⁹ Panel (b) reveals the calibration curve for determining the Ni^{2+} concentrations at equilibrium.

Adsorption capacity

The adsorption capacity (mg g^{-1}) was estimated by using the following eqn (1):

$$q_e = \frac{(C_i - C_e) \text{ mg L}^{-1}}{W_g} \times V_L \quad (1)$$

where: C_i is the initial concentration of Ni^{2+} (mg L^{-1}), C_e is the concentration at equilibrium (mg L^{-1}), W_g is the mass of GO (g), V_L is the volume of sample.

The adsorption capacity (q_e , mg g^{-1}) at different time intervals is increased with increasing the initial Ni^{2+} concentrations as shown in Fig. 4. Moreover, the lateral width of GO particles exhibits a major change in the adsorption capacity due to more surfaces of GO – 200 nm are available. As shown in Fig. 4a, the adsorption capacity of GO – 450 nm appeared as stabilized behavior above 400 mg L^{-1} of Ni^{2+} concentration. In panel b, the GO – 200 nm exhibits a linear behavior with increasing Ni^{2+} concentration up to 900 mg L^{-1} after 60 min of incubation.

Percent removal of Ni^{2+} at different concentrations and time intervals

Fig. 3 shows the percent removal of Ni^{2+} ions, it depends on the lateral width of GO particles. As the size of GO decreases the percent removal increases due to high surface areas of GO – 200 nm. As shown in panel (a) the contact time reveals a minor change on the removal of Ni^{2+} after 60 min of incubation. In panel (b) the size of GO particles (GO – 200 nm) shows higher removal of Ni^{2+} due to increasing adsorbate surfaces.^{1,24}

The removal of Ni^{2+} percent was calculated by using the following eqn (2):

$$\% R = \frac{C_i - C_e}{C_i} \times 100 \quad (2)$$

where: C_i (mg L^{-1}) is the initial concentration of Ni^{2+} , C_e (mg L^{-1}) is the concentration at equilibrium deduced from the calibration curve.⁴⁰

Percent removal of Ni^{2+} at different pH

The percentage removal of Ni^{2+} at different pH values from 2 to 10 is shown in Fig. 5, this figure shows that the Ni^{2+} uptake was increased between 2 and 6 pH values. This indicates that the competitive adsorption between Ni^{2+} ions and H_3O^+ ions on the surface of GO. At a pH higher than 6, the negative charge on the surface of GO particles increases and lead to strong electrostatic interaction between Ni^{2+} and GO particles under alkaline condition.^{18,24} On the other hand, Ni^{2+} ions can present as $(\text{Ni}(\text{OH})^+)$, $\text{Ni}(\text{OH})_2$, $\text{Ni}(\text{OH})_3^-$, $\text{Ni}(\text{OH})_4^{2-}$ and may be precipitated onto the surface of GO particles.^{1,38}

The lateral sheet dimension (GO – 200 nm) shows higher removal comparing with GO – 450 nm. This behavior is attributed to more surfaces of GO particles. In conclusion, the

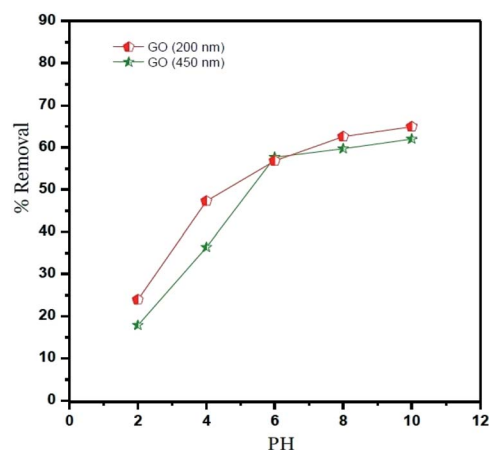


Fig. 5 Percent removal of Ni^{2+} at different pH values. $\text{Ni}^{2+} = 400 \text{ mg L}^{-1}$, $T = 25^\circ\text{C}$.

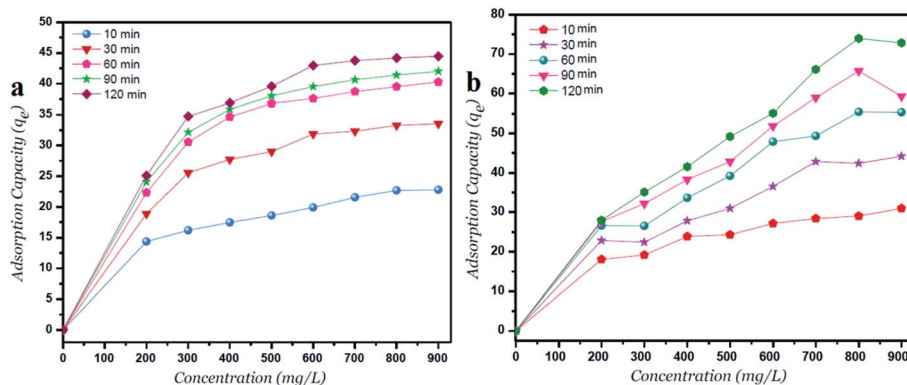


Fig. 4 Adsorption capacity versus Ni^{2+} concentrations for (a) GO – 450 nm and (b) GO – 200 nm at different time intervals.



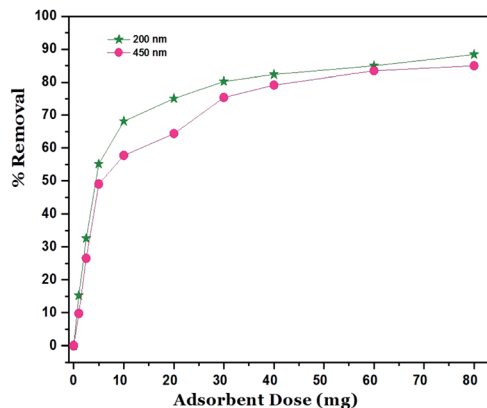


Fig. 6 Effect of adsorbent dose ($\text{Ni}^{2+} = 400 \text{ mg l}^{-1}$, $T = 25^\circ\text{C}$, $\text{pH} = 6$).

optimum pH value of Ni^{2+} adsorbed onto GO – 450 and GO – 200 nm is 6, which was used for further studies.

Adsorbent dose

The effect of adsorbent dose on Ni^{2+} removals is shown in Fig. 6. The amount of GO varying from 1.0 to 80 mg and the initial Ni^{2+} concentration for adsorbent dose was fixed at 400 mg l^{-1} . The results reveal that the removal of Ni^{2+} by using GO – 450 nm and GO – 200 nm increased with an increasing of adsorption dose from 1.0 mg to 20 mg, this behaviour is attributed to the increasing of the availability of surfaces at a higher amount of adsorbents. A plateau was reached from 20 up to 80 mg of adsorbent, and this may be imputed to overcrowding in the adsorbed molecules.¹

Adsorption isotherm models

The adsorption isotherm plays a key role for the analysis of adsorption process that elucidates how the Ni^{2+} ions distribute between the solid-liquid interface once the system is at equilibrium. The most model used is Langmuir isotherm for monolayer adsorption onto the adsorbent particles. It can be expressed as eqn (3).

$$\frac{C_e}{q_e} = \frac{C_e}{q_{\max}} + \frac{1}{K_L q_{\max}} \quad (3)$$

where: q_e is the amount of Ni^{2+} ions adsorbed per unit weight of adsorbent (mg g^{-1}), C_e is the Ni^{2+} concentration at equilibrium (mg l^{-1}). q_{\max} is the maximum sorption capacity specified by the reactive species in an ideal monolayer system (mg g^{-1}). K_L is the Langmuir constant affined to a free energy of adsorption (l mg^{-1}).

The Langmuir parameter (R_L) is a dimensionless constant separation factor defined as in eqn (4)

$$R_L = \frac{1}{1 + K_L C_i} \quad (4)$$

where: C_i is the maximum initial concentration of adsorbate (mg l^{-1}), K_L is the Langmuir constant (l mg^{-1}). R_L value indicates the adsorption reliable if $R_L > 1$ unfavorable, $0 < R_L < 1$ favorable, $R_L = 1$ for linear adsorption, and $R_L = 0$ for irreversible adsorption.¹

The Freundlich model is an empirical equation consider that adsorption of metal ions occurs on heterogeneous surface by multilayer adsorption; the linear form of Freundlich can be expressed in eqn (5).

$$\ln q_e = \ln K_f + \frac{1}{n} \ln C_e \quad (5)$$

where: K_f and (n) are the Freundlich parameters related to adsorption capacity and adsorption intensity respectively.⁴¹

Langmuir and Freundlich isotherm models using GO – 450 nm

A plot of C_e/q_e versus C_e at different time intervals gives a linear relation with the slope of $1/q_{\max}$ and intercept of $1/(K_L q_{\max})$ as shown in Fig. 7a. The Langmuir model postulates that the uptake of Ni^{2+} ions occurs on a homogenous surface by monolayer adsorption without interaction between the adsorbed ions. For data collection at equilibrium, initial metal concentrations were varied while the GO adsorbent mass remained constant. The adsorption parameters of the Langmuir isotherm model are shown in Table 1 (ESI†) at different time intervals. The correlation coefficient of Langmuir isotherm (R^2) was approximately 0.999, which reveals a high correlation or linear relationship. Therefore the adsorption of Ni^{2+} onto GO – 450 nm is correlated well with the Langmuir equation.

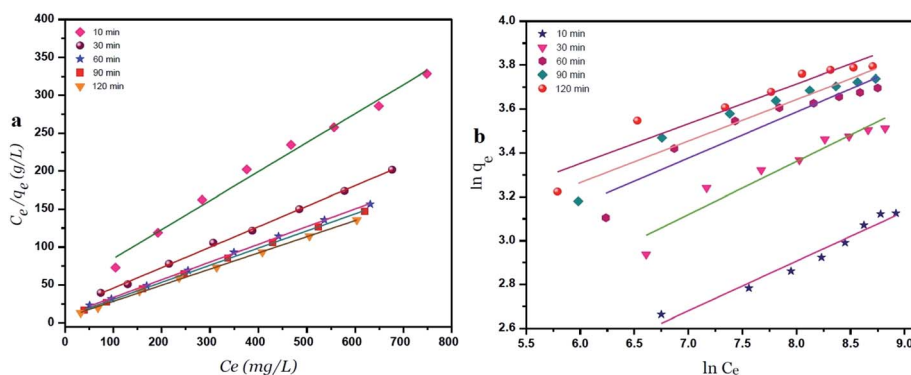


Fig. 7 Langmuir (a) and Freundlich (b) adsorption isotherm of Ni^{2+} onto GO – 450 nm at 25°C .



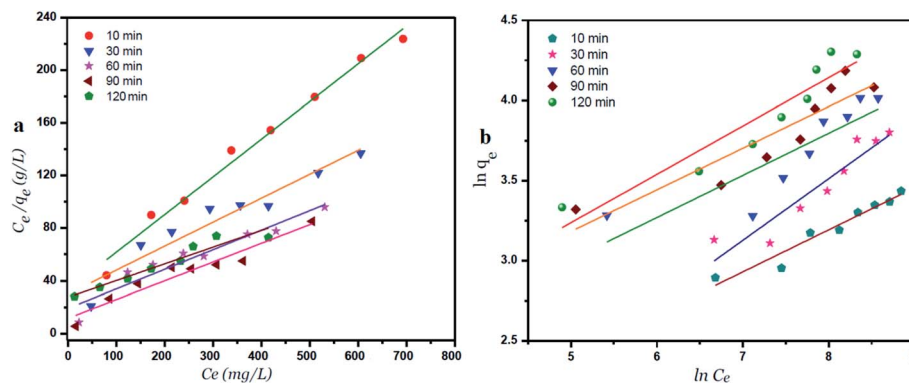


Fig. 8 Langmuir (a) and Freundlich (b) adsorption isotherm of Ni^{2+} onto GO – 200 nm at 25 °C.

The Freundlich isotherm model interpreted as sorption to heterogeneous surfaces as shown in Fig. 7b. The Freundlich constant n between 1 and 10 provides favourable adsorption tends. The larger value of n means stronger interaction between the Ni^{2+} ions and GO nanoparticles as shown in Table 1 (ESI†). The Freundlich isotherm parameter (K_F) was calculated according to eqn (5), the value increases as the time interval increased which indicates the Freundlich model fitted well with the experimental data. The correlation coefficient of Freundlich isotherm (R^2) was approximately 0.910.

Langmuir and Freundlich isotherm models using GO – 200 nm

Fig. 8 exhibits the Langmuir and Freundlich adsorption isotherm of Ni^{2+} onto GO – 200 nm at 25 °C. Panel (a) shows the linear relation between C_e/q_e versus Ni^{2+} concentrations at equilibrium for different time intervals. The values of K_L and q_m calculated from the plotted data (slope and intercept) are reported in Table 2 (ESI†). The value of q_m of Ni^{2+} onto GO – 200 nm is higher than that onto GO – 450 nm. This result indicated that a complete and uniform monolayer of Ni^{2+} covering the surfaces of GO particles over the whole concentrations. However the R^2 values at different time intervals were less than that onto GO – 450 nm, indicating that the size of GO – 200 nm particles affect Ni^{2+} uptake onto the adsorbent.

Langmuir model indicated the best fitting ($R^2 = 0.999$) to the isotherm data is with GO – 450 nm size distributions.

In panel (b), the same linear relation appeared between $\ln q_e$ versus $\ln C_e$ for Freundlich isotherm. The value of n is higher than 1 indicated that the adsorption capacity was slightly restrained at lower equilibrium concentration. As shown in Table 2 (ESI†), the values of n are studied as an indication of the linearity deviation, as well as it is used to predict the heterogeneity degree of the adsorbent. Moreover the value of n is considered as an indication whether the adsorption process is favorable or not. The value of n corresponding to GO – 450 nm is greater than that with GO – 200 nm, illustrating that a stronger capacity of Ni^{2+} onto GO – 450 nm. The Freundlich constant K_F has been found as a relative measure of adsorption capacity. As the value of K_F increases the adsorption capacity increases. The reported values of K_F by using GO – 450 nm as adsorbent were greater than that using GO – 200 nm. This result summarizing that the uptake of Ni^{2+} ions with high adsorptive capacity of adsorbent.

In this work, the data of Ni^{2+} onto GO – 450 nm and GO – 200 nm of equilibrium adsorption was better fitted to Langmuir than Freundlich adsorption isotherm model. The maximum adsorption capacities of monolayer were 43.478 mg g^{-1} and 66.667 mg g^{-1} for GO – 450 nm and GO – 200 nm respectively as shown in Fig. 9.

The isotherm parameters are summarized in Table 1.

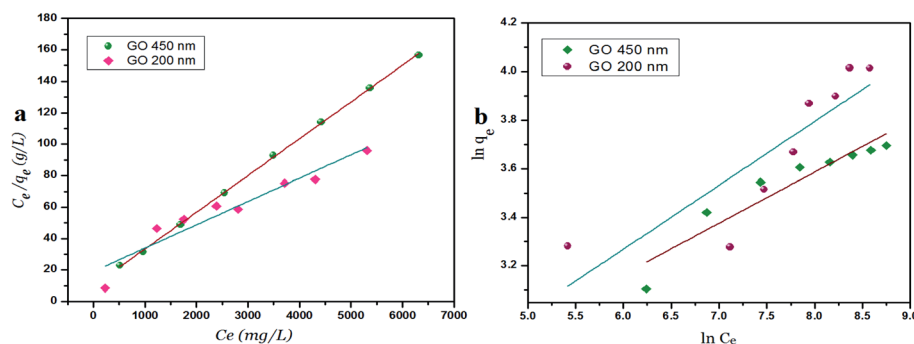
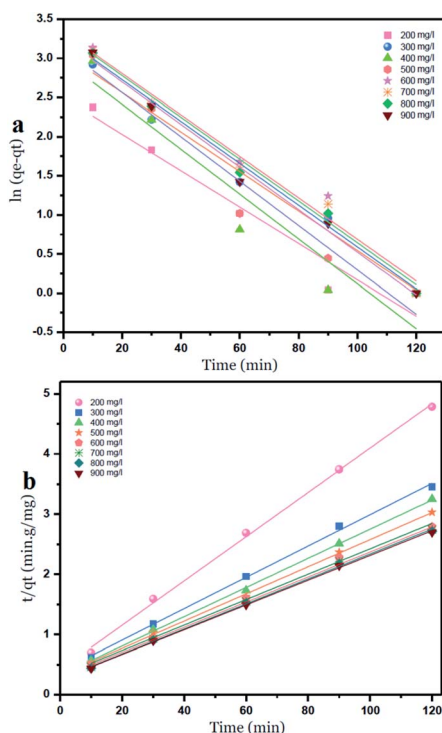


Fig. 9 Langmuir (a) and Freundlich (b) adsorption isotherm of Ni^{2+} onto 450 nm and 200 nm at 60 min of incubation.



Table 1 Langmuir and Freundlich isotherm parameters

Adsorbents	Langmuir			Freundlich	
	q_m	k_l	R_L	n	K_f
GO 450 nm	43.478	0.0023	0.046	4.762	6.686
GO 200 nm	66.667	0.0008	0.124	3.802	5.419

Fig. 10 Pseudo first order (a) and Pseudo second order kinetic adsorption model of Ni^{2+} onto GO – 450 nm at 25 °C.

Kinetic study

Two kinetic models were used to investigate the adsorption of Ni^{2+} ions onto GO particles. The models were considered as follows:

Lagergren's pseudo-first-order model which can be expressed by eqn (6):

$$\ln q_e - q_t = \ln q_e - k_1 t \quad (6)$$

where: q_e is the adsorption capacity at equilibrium (mg g^{-1}), q_t is the adsorption capacity at any time t (min). k_1 is the first-order rate constant adsorption min^{-1} . Values of q_e and k_1 were calculated by plotting of $\ln q_e - q_t$ versus time.

Ho's pseudo-second-order model is given by eqn (7):

$$\frac{t}{q_t} = \frac{t}{q_e} + \frac{1}{k_2 q_e^2} \quad (7)$$

where: k_2 is the rate constant for the pseudo-second-order adsorption ($\text{g mg}^{-1} \text{min}^{-1}$), values of k_2 and q_e for different initial concentrations of ions were calculated from the intercept and slope respectively.⁴²

The intercepts and slope of the linear plots for the adsorption of nickel by GO – 450 nm and GO – 200 nm were used to calculate the kinetic parameters, as shown in Fig. 10 and 11, the kinetic constant values are summarized in Table 6 (ESI†).

Effect of temperature

The adsorption of Ni^{2+} ions was studied at different temperatures (25 °C, 45 °C, and 65 °C), the adsorption data were carried out with 400 mg l^{-1} of Ni^{2+} at pH = 6.

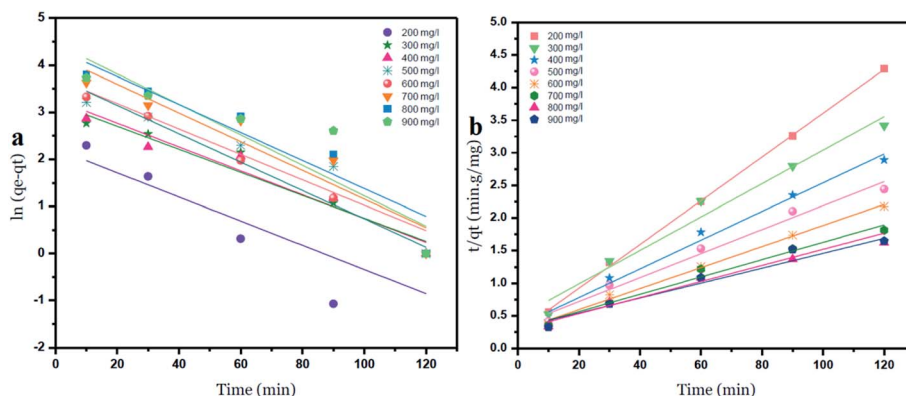
The parameters of thermodynamic like entropy ΔS° , enthalpy ΔH° , and Gibbs free energy ΔG can be expressed by Van't Hoff equation

$$\ln \left(\frac{q_e}{C_e} \right) = \frac{\Delta S^\circ}{R} - \frac{\Delta H^\circ}{R \cdot T} \quad (8)$$

where: T (K) is the solution temperature. R is the gas constant ($8.314 \text{ J mol}^{-1} \text{ K}^{-1}$)

From the plotting of $\ln(q_e/C_e)$ versus $(1/T)$ entropy (ΔS°) and enthalpy (ΔH°) were calculated by using the intercept and slope respectively as shown in Fig. 12, and the Gibbs free energy was calculated by using this eqn (8).^{18,43}

As shown in Table 9 (ESI†), the calculated values of the thermodynamic parameters are elucidated. The positive values of ΔH° indicate that the adsorption process of Ni^{2+} onto GO – 450 nm and 200 nm is an endothermic process while the

Fig. 11 Pseudo first order (a) and Pseudo second order (b) kinetic adsorption model of Ni^{2+} onto GO – 200 nm at 25 °C.

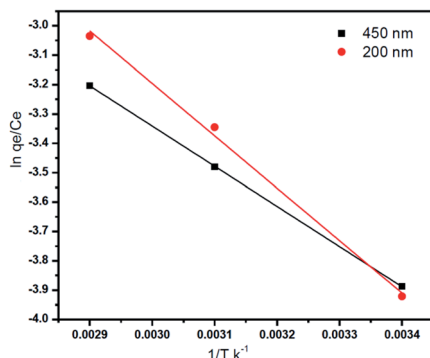


Fig. 12 Van't Hoff for the adsorption of Ni²⁺ onto GO – 450 nm and GO – 200 nm. (Ni²⁺ = 400 mg l⁻¹).

positive values of ΔS° indicate that increased the randomness at the solid/liquid interface during the adsorption process. The negative values of ΔG° suggest that the adsorption of Ni²⁺ onto GO – 450 nm and 200 nm is a spontaneous process but as these values are in the range between -20 and 0 kJ mol⁻¹ the process is classified physisorption.^{1,18}

Surface area measurements of GO materials

The specific surface area of graphene oxide (450 nm, 200 nm) was calculated by the following equation:

$$S = \frac{q_m \times a_{MB} \times N_A \times 10^{-20}}{M} \quad (9)$$

where: S is the specific surface area in m² g⁻¹; q_m is the maximum adsorption capacity of methylene blue at the monolayer of GO in mg g⁻¹, a_{MB} is the occupied surface area of one molecule of methylene blue = 130 Å², N_A represents Avogadro's number; and M is the molecular weight of methylene blue, 373.9 g mol⁻¹. The maximum adsorption capacity (q_m) of GO – 450 and GO – 200 nm is 357.1 and 502.5 mg g⁻¹ respectively which was evaluated from eqn (3).

Conclusion

The removal of heavy metal ions from aqueous solution can be achieved by using graphene oxide particles at different size distributions (GO – 450 nm and GO – 200 nm). The smaller size of GO particles provided better removal due to high surface to volume ratio, as well as various oxygen groups like hydroxyl, epoxy and carboxyl present after oxidation and ultrasonication. The adsorption capacity of GO – 450 nm appeared as stabilized behavior above 400 mg l⁻¹ Ni²⁺ concentration. However the GO – 200 nm exhibits a linear behavior with increasing Ni²⁺ concentration upto 900 mg l⁻¹ after 60 min of incubation. This research demonstrates that GO particles can be an effective adsorbent for toxic metal removal depending on the particle's size distribution.

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Conflicts of interest

The authors declare that no competing interests.

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