



Charge transfer from and to manganese phthalocyanine: bulk materials and interfaces

Florian Ruckerl¹, Daniel Waas¹, Bernd Büchner¹, Martin Knupfer^{*1}, Dietrich R. T. Zahn², Francisc Haidu², Torsten Hahn³ and Jens Kortus³

Review

[Open Access](#)

Address:

¹IFW Dresden, Helmholtzstr. 20, D-01069 Dresden, Germany,
²Semiconductor Physics, Chemnitz University of Technology,
D-09107 Chemnitz, Germany and ³Institute of Theoretical Physics,
TU Bergakademie Freiberg, Leipziger Str. 23, D-09596 Freiberg,
Germany

Email:

Martin Knupfer* - m.knupfer@ifw-dresden.de

* Corresponding author

Keywords:

charge transfer; electronic properties; manganese phthalocyanine

Beilstein J. Nanotechnol. **2017**, *8*, 1601–1615.

doi:10.3762/bjnano.8.160

Received: 03 March 2017

Accepted: 13 June 2017

Published: 04 August 2017

This article is part of the Thematic Series "Towards molecular spintronics".

Guest Editor: G. Salvan

© 2017 Ruckerl et al.; licensee Beilstein-Institut.

License and terms: see end of document.

Abstract

Manganese phthalocyanine (MnPc) is a member of the family of transition-metal phthalocyanines, which combines interesting electronic behavior in the fields of organic and molecular electronics with local magnetic moments. MnPc is characterized by hybrid states between the Mn 3d orbitals and the π orbitals of the ligand very close to the Fermi level. This causes particular physical properties, different from those of the other phthalocyanines, such as a rather small ionization potential, a small band gap and a large electron affinity. These can be exploited to prepare particular compounds and interfaces with appropriate partners, which are characterized by a charge transfer from or to MnPc. We summarize recent spectroscopic and theoretical results that have been achieved in this regard.

Review

Introduction

The family of metal-centered phthalocyanines has been considered for future technological applications because of their favorable electronic and optical properties and their advantageous chemical stability [1-8]. Phthalocyanine molecules can harbor a number of metal ions, in particular transition-metal ions such as cobalt, iron or manganese. A special characteristic of transition-

metal centered phthalocyanines is, that transition-metal ions often are characterized by a magnetic moment, and therefore such phthalocyanines also show very interesting magnetic behavior [9]. They have even been discussed in terms of molecular magnets including their discussion in future applications in the field of molecular spintronics [10-12].

Among these transition-metal phthalocyanines, manganese phthalocyanine (MnPc) is one of the most interesting molecules due to its particular electronic and magnetic properties in the bulk [13–16]. A schematic representation of the molecule structure is depicted in Figure 1 below. For instance, MnPc is characterized by an unusual $S = 3/2$ spin state of the central Mn(II) ion. The spin of MnPc is a consequence of three unpaired 3d electrons in the Mn 3d levels, which also lie close to the chemical potential. In essence, the electronic properties (partly) reflect the behavior of these 3d electrons, and MnPc plays a special role in the group of the transition-metal phthalocyanines. The energy gap between the occupied and unoccupied molecular orbitals of MnPc is the smallest among all transition-metal phthalocyanines [17–21], its ionization potential also is the smallest within this class of material [17,18,22], while the electron affinity is larger than those of the others [17,18]. Furthermore, the exciton binding energy related to the lowest electronic singlet excitation is somewhat larger compared to, e.g., CuPc [18–20,23]. In Table 1 we summarize these values in comparison to CuPc, the most prominent and most extensively investigated transition-metal phthalocyanine to date.

Table 1: Summary of characteristic electronic parameters for MnPc determined for thin films in comparison to those for CuPc. In detail, we compare the ionization potential (IP), the electron affinity (EA), the energy gap as seen in optical studies (E_g^{opt}), the transport energy gap (E_g^{tr}), and the exciton binding energy (E_B^{exc}) of the lowest singlet excitation (see text for references). All values are given in eV.

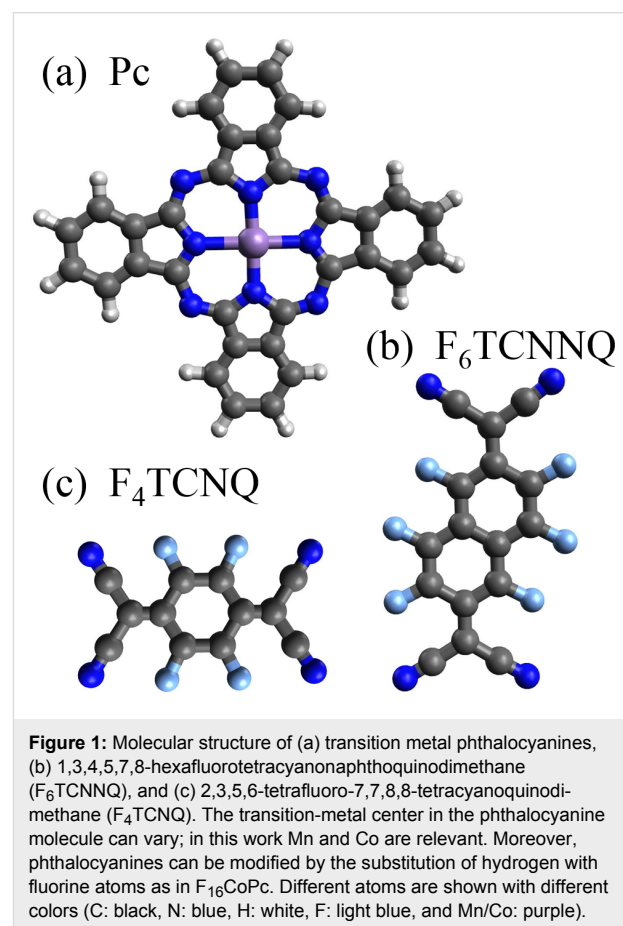
	IP	EA	E_g^{opt}	E_g^{tr}	E_B^{exc}
MnPc	4.5	3.3	0.5	1.2	0.7
CuPc	5.0	2.7	1.8	2.3	0.5

These exceptional properties of MnPc render it possible that this molecule can undergo charge-transfer reactions of either kind, i.e., it can be oxidized or reduced by suitable reaction partners. This can be utilized to synthesize new compounds with potentially interesting properties. In this contribution we present a summary of recent results in regard of charge transfer compounds, or interfaces characterized by charge transfer, which all are based on MnPc.

Materials and methodology

This article covers charge-transfer reactions of manganese phthalocyanine with the alkali metal potassium as well as with further organic molecules. The latter are characterized by a rather large electron affinity to enable charge transfer from MnPc to these structures. In Figure 1 we show the structure of all molecules discussed below. These are MnPc, its phthalocyanine relative $F_{16}\text{CoPc}$ (both purchased from Sigma-Aldrich),

$F_4\text{TCNQ}$ (2,3,5,6-tetrafluoro-7,7,8,8-tetracyanoquinodimethane, TCI Europe) and $F_6\text{TCNNQ}$ (1,3,4,5,7,8-hexafluorotetracyanonaphthoquinodimethane, Novaled AG). $F_4\text{TCNQ}$ is quite well known for its high electron affinity [24], and it is also known to form many charge-transfer crystals with appropriate partners [25,26]. Further, it has also been used as dopant material for organic electronic devices [27–29]. More recently, $F_6\text{TCNNQ}$ has been introduced into organic devices with advantages such as an even higher electron affinity and a larger molecular mass, which prevents diffusion of the molecule in organic devices [30]. $F_{16}\text{CoPc}$ has been applied recently in a few cases only, in order to induce charge transfer across interfaces to other insulating (or semiconducting) materials [31–33]. The electron affinities of the three acceptor molecules are 4.5 eV ($F_{16}\text{CoPc}$ [34]), 5.2 eV ($F_4\text{TCNQ}$ [35]) and 5.6 eV ($F_6\text{TCNNQ}$ [36]).



The results presented in this contribution were achieved by either solid-state spectroscopy methods or density functional based calculations. The experimental methods comprise photoelectron (or photoemission) spectroscopy (PES), inverse photoemission spectroscopy (IPES), electron energy-loss spectroscopy (EELS), spectroscopic ellipsometry and X-ray absorption

spectroscopy (XAS). Here, we only briefly mention the kind of information that is provided by these methods, and we refer the reader to comprehensive literature for detailed information.

PES [37-39] is based on the photoelectric effect and provides insight into the valence-band electronic density of states as well as the binding energy and line shape of core levels, which give information about the composition of the sample and the chemical state (e.g., valency) of the atoms or ions. In IPES [19,40-42], the unoccupied density of states is probed. EELS [43-45] can also be called inelastic electron scattering and measures the electronic excitations either in the valence-band region, or from core levels into unoccupied states, whereas momentum-dependent studies are possible [43,45,46]. The EELS cross section is proportional to $\text{Im}(-1/\epsilon)$ (ϵ is the dielectric function). In this way, one can investigate valence-band excitations (cf. optical methods) and the element-projected unoccupied density of states. Also, access to orbital selective occupations and the magnetic moment of open shells is accessible. Spectroscopic ellipsometry [47-49] measures the change in the light polarization after reflection on a sample surface. This information allows for the determination of the real and the imaginary part of the dielectric function. XAS [42,50] is equivalent to EELS in the core-level region, and polarization-dependent studies have often been carried out to study the molecular orientation on substrates. In addition to our experiments we performed calculations within the density functional theory (DFT) framework. We used a recent version of the NRLMOL all-electron DFT code [51,52], which uses large Gaussian-orbital basis sets for the representation of the electronic wavefunctions [53]. Unless noted otherwise we used the PBE functional [54] within the general gradient approximation (GGA) was used for all calculations. We semi-empirically included dispersion correction according to the Grimme DFT-D3 method [55] in all of our calculations.

$K_x\text{MnPc}$: formation of stable phases with MnPc anions

The formation of compounds with composition $K_x\text{MnPc}$ was achieved by evaporation of potassium from so-called K dispensers (SAES Getters, S.P.A, Italy) onto MnPc thin films under ultra-high vacuum conditions. These then were thoroughly characterized by the spectroscopic methods in order to obtain a comprehensive picture. We start the presentation of our spectroscopic data with the development of the electronic excitation spectra of MnPc as a function of potassium doping. These were determined using EELS and spectroscopic ellipsometry. In Figure 2 we show the corresponding results.

The EELS measurements were carried out at a momentum transfer of 0.1 \AA^{-1} which represents the so-called optical limit,

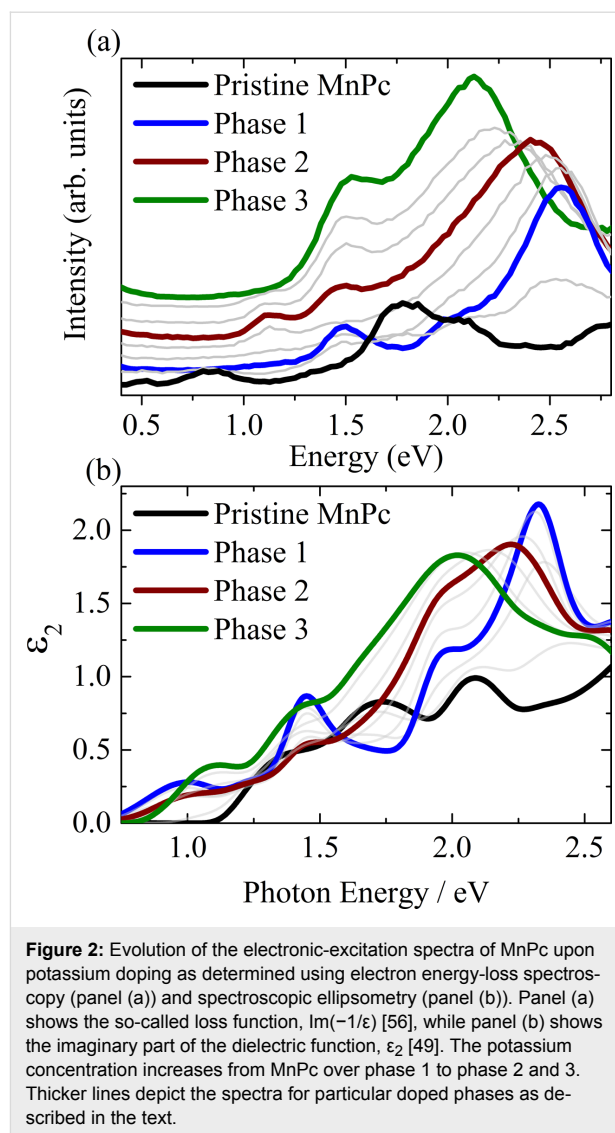


Figure 2: Evolution of the electronic-excitation spectra of MnPc upon potassium doping as determined using electron energy-loss spectroscopy (panel (a)) and spectroscopic ellipsometry (panel (b)). Panel (a) shows the so-called loss function, $\text{Im}(-1/\epsilon)$ [56], while panel (b) shows the imaginary part of the dielectric function, ϵ_2 [49]. The potassium concentration increases from MnPc over phase 1 to phase 2 and 3. Thicker lines depict the spectra for particular doped phases as described in the text.

i.e., the data are equivalent to those from corresponding optical studies [43]. Note that the ellipsometry results in Figure 2b start at higher energies due to instrumental limitations. The spectrum of pristine MnPc is characterized by several spectral features at about 0.5, 0.8, 1.4, 1.8, and 2.1 eV [23], which are clearly seen in Figure 2a. Around 2 eV, the excitations are usually ascribed to the Q band common to many phthalocyanines [57-59]. They are due to transitions from the highest occupied molecular orbital (HOMO) to the lowest unoccupied molecular orbital (LUMO) [58]. The observed fine structure was attributed to a combination of Davydov splitting, the admixture of charge-transfer excitations as well as vibronic satellites [60-68]. The appearance of further excitations at lower energies for MnPc is still not fully understood. Certainly, they are connected to the contribution of Mn 3d states to the molecular orbitals close to the chemical potential, a complete picture, however, is elusive [23,69-73].

The addition of potassium induces particular changes in regard of the electronic excitations. The two lowest-lying features (at about 0.5 and 0.8 eV, see panel (a)) disappear, also the intensity in the Q band region is drastically reduced. Instead, spectral structures show up at 1.5 eV and about 2.6 eV up to a particular doping level (called phase 1 in Figure 2). Further potassium doping results in the appearance of an excitation at about 1.1 eV, while the structure at about 2.6 eV shows a downshift in energy. The intensity of the excitation at 1.1 eV as seen in panel (a) reaches a maximum at a second distinct doping level (phase 2), thereafter it starts to vanish again. In contrast, the excitation at 1.5 eV is steadily growing in intensity and the highest-lying excitation continues to downshift until a third composition is reached (phase 3).

A detailed analysis of these data together with those from electron diffraction [56] revealed the existence of particular $K_x\text{MnPc}$ compositions (phases 1, 2, and 3). This, in general, parallels the behavior of other transition-metal phthalocyanines upon potassium doping, where also particular stable phases were reported [74–76]. We emphasize that this conclusion is nicely supported by the fact that all EELS spectra at doping levels between these three phases can be modeled by a corresponding superposition of the spectra of the phases in the direct neighborhood [56]. The exact composition of these phases was finally determined by an analysis of the respective C 1s and K 2p core-level excitations. These data are depicted in Figure 3.

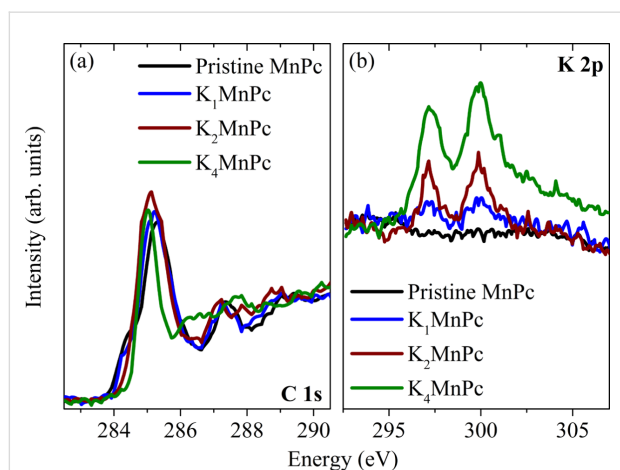


Figure 3: C 1s (panel (a)) and K 2p (panel (b)) excitation edges of MnPc and the three potassium-doped phases (adapted from [56]). The relative intensities of these two edges were used to analyze the composition of the doped phases.

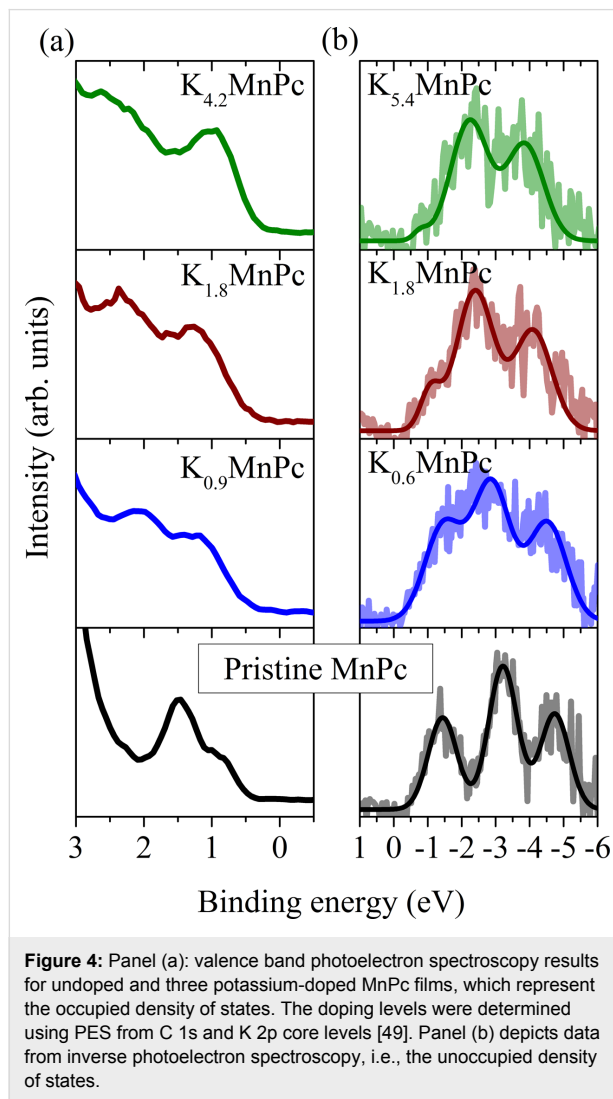
In this Figure, excitation from the C 1s core level in unoccupied π -derived states start at about 285 eV, those into σ -derived carbon states at about 292 eV, while K 2p to K 3d excitations are seen at 297.1 and 298.8 eV [50,77]. The spectral evolution

in Figure 3 clearly signals the increasing K content in our samples. The analysis of the relative spectral weights allowed for the determination of the exact composition of the three potassium-doped MnPc phases: $K_1\text{MnPc}$, $K_2\text{MnPc}$, and $K_4\text{MnPc}$ [56]. Thus, the spectra highlighted in Figure 2 above represent the electronic excitation spectra of these three phases. In addition, the spectral shape of the C 1s excitation data change in a characteristic manner as a function of doping, in particular right at the excitation onset.

In the case of undoped MnPc, the excitation edge starts with a low-energy shoulder around 284.5 eV before the first maximum at 285.3 eV is reached. Also for other transition-metal phthalocyanines such a C 1s excitation edge is observed [78–80]. This two-peak structure has its origin in the fact that the phthalocyanine ring consists of carbon atoms with different chemical environments, namely those with nitrogen as neighbors and those without. This is clearly seen in corresponding C 1s core-level photoemission data [74,81,82], in which the binding energy for the two carbon species is different. Considering the C 1s excitations as seen in Figure 3a, the excitations from these two carbon sites into the lowest unoccupied orbital give rise to a two-peak structure. Upon doping, there are two changes that impact the C 1s excitations. First, the unoccupied levels are filled with the doping-induced electrons. Second, the carbon binding energies change, as revealed by photoemission data, which show a broadening and the appearance of significantly less structured C 1s core level features in the doped compounds [74,79,80]. As a consequence, the low-energy shoulder in the C 1s excitations is lost and only a single low-energy feature is seen for higher doping levels. Again, this parallels the observations for other doped phthalocyanines [78–80]. There is, however, an important difference to the evolution of the C 1s excitation edges of FePc, CoPc and CuPc upon potassium doping. For these three materials the low-energy shoulder virtually disappeared at a doping level of about one K atom per molecule [78,79,83]. In the case of MnPc here (see Figure 3), this state is reached not until the composition $K_2\text{MnPc}$ is reached. This nicely corroborates that the lowest-lying unoccupied MnPc orbital that is filled by electrons, has predominantly Mn 3d character [22,72,84], whereas for the other phthalocyanines a ligand π^* state is filled right from the beginning.

The results shown above already indicate that all of the potassium-doped MnPc phases are characterized by an energy gap. This observation is in full agreement with the results of photoelectron spectroscopy studies and inverse photoelectron studies of K-doped MnPc [49] as discussed in the following. The samples for these investigations again were prepared by potassium addition to MnPc thin films in ultra-high vacuum. The doping level of the films was determined by analyzing the relative in-

tensities of the core level photoemission from the C 1s and the K 2p core levels [49]. In Figure 4 we summarize the results of PES and IPES data that could be obtained for compositions close to the K_x MnPc phases that were discussed above.



We start the discussion with the spectra for undoped MnPc. The PES data in Figure 4a show the well-known two maxima close to the chemical potential (0 eV binding energy) [18,22]. They arise from emission from the highest occupied molecular orbital (HOMO) at about 1.5 eV binding energy, and from the so-called SOMO (singly occupied molecular orbital) at about 0.7 eV. Going to $K_{0.9}$ MnPc, there is an energy shift to higher binding energies, which is due to a shift of the Fermi level towards the unoccupied levels. Furthermore, the feature at lowest binding energy grows in intensity, which reflects the filling of the SOMO with one more electron. Adding more potassium ($K_{1.8}$ MnPc) results in a further shift of the maxima to somewhat higher energies, while the spectral onset moves

slightly downward. This downshift is related to the filling of the former LUMO with electrons, i.e., a new occupied state appears in the photoemission data. Further doping to $K_{4.2}$ MnPc causes an intensity increase at low binding energy, which results from further filling of the former LUMO.

Figure 4b presents information on the unoccupied electronic states as measured using IPES [49]. For undoped MnPc the spectrum represents the first three unoccupied levels, the LUMO (lowest unoccupied molecular orbital), the LUMO+1, and the LUMO+2. We note that the IPES data do not reveal the unoccupied part of the SOMO (see above), which is attributed to the limited energy resolution of the data. The formation of K_x MnPc phases results in an intensity decrease of the feature representing the LUMO due to electron addition into this orbital, in agreement to the discussion above. Importantly, independent of the potassium concentration both PES and IPES data, in agreement with the EELS data above demonstrate a clear energy gap, i.e., none of the K_x MnPc phases is metallic. This, in general, resembles the situation in many molecular crystals doped with alkali metals, where it was observed that the doping did not result in a metallic ground state although metallicity would be expected on the basis of band-structure calculations since half-filled bands are present. Molecular crystals usually have energy bands with small band widths, which is a direct consequence of the rather small interaction between the molecules in the material. Furthermore, the bandwidth often is similar to the Coulomb repulsion of two charge carriers on one molecule. Thus, molecular crystals also are correlated materials, where the electronic correlations often are strong enough to induce an insulating Mott–Hubbard ground state [85-94].

MnPc cations in an organic salt: MnPc/F₄TCNQ

In the following we present information on the electronic properties of a purely organic salt in which MnPc is oxidized. In order to obtain such a compound we have prepared mixed films consisting of MnPc and the particularly strong electron acceptor F₄TCNQ. It is well known that high-quality organic charge transfer crystals with F₄TCNQ as electron acceptor can be grown [95-103]. Phthalocyanine-based compounds, however, have not been reported yet. We prepared the mixed films by evaporating F₄TCNQ on top of a MnPc film and taking advantage of the diffusion of F₄TCNQ into the MnPc film, or by simultaneously evaporating both materials [104]. Subsequently, we have applied an in situ distillation procedure, which was already applied in previous experiments to achieve stoichiometric molecular salts [105-108]. After the initial preparation step, the films were heated up to 340 K for about half an hour, where an evaporation of surplus F₄TCNQ could be seen even with bare eye [104]. Also, this procedure resulted in the forma-

tion of MnPc/F₄TCNQ films with a well-defined composition and with well-defined spectral response, which signalled the homogeneity of the films. The composition was analyzed using the relative intensities of photoemission core-level features [104], and in all cases we obtained a stoichiometric ratio of 1:1.

We have analyzed the electronic properties of the new organic charge-transfer compound MnPc/F₄TCNQ using photoemission and electron energy-loss spectroscopy as well as density functional theory based calculations. In Figure 5 we show the PES results in the valence band region.

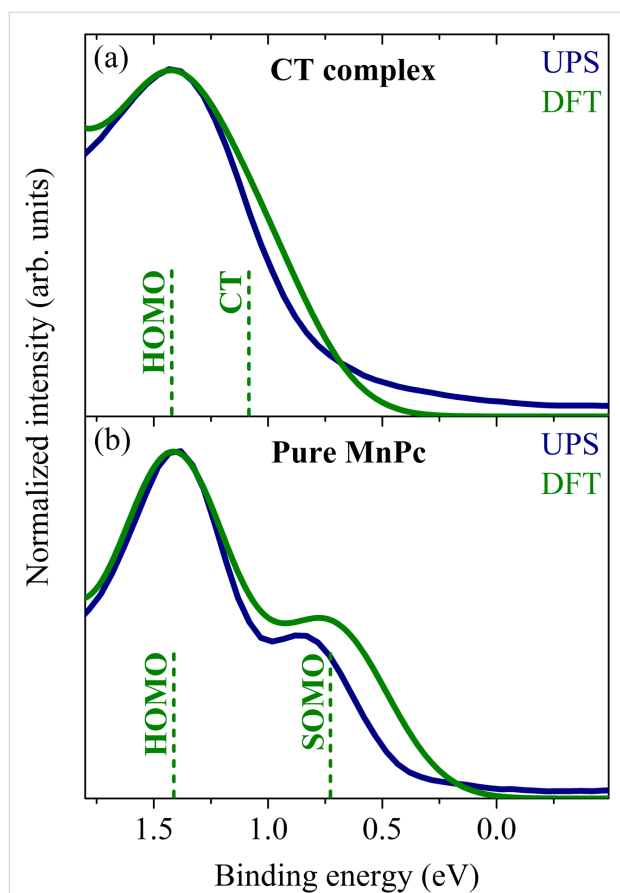


Figure 5: Comparison of valence band photoelectron spectroscopy (UPS) data to those from density functional based calculations (DFT) for pure MnPc (lower panel) and the charge-transfer compound MnPc/F₄TCNQ (upper panel). The vertical bars denote the energy position of the molecular orbitals as determined by the calculations. For pure MnPc the two-peak structure arises from the singly occupied molecular orbital (SOMO) and the highest occupied molecular orbital (HOMO). The formation of the charge-transfer (CT) compound leads to a disappearance of the SOMO. Instead, a CT hybrid state shows up, closer in energy to the HOMO (see also [104]).

These data clearly illustrate a substantial variation going from pure MnPc to the charge-transfer compound. For pure MnPc our data demonstrate the well-known two peak structure at lowest binding energy (see also the previous chapter). The for-

mation of the MnPc/F₄TCNQ compound is accompanied by the disappearance of the feature at lowest binding energy (here called SOMO = singly occupied molecular orbital). This indicates that electrons are removed from MnPc, i.e., the phthalocyanine molecule is oxidized. Since the leading orbital in pure MnPc (SOMO) is a hybrid state of Mn 3d and ligand π orbitals with a large Mn 3d contribution [22,84,109], this oxidation is also clearly seen in core-level photoemission from MnPc 2p core level states as depicted in Figure 6.

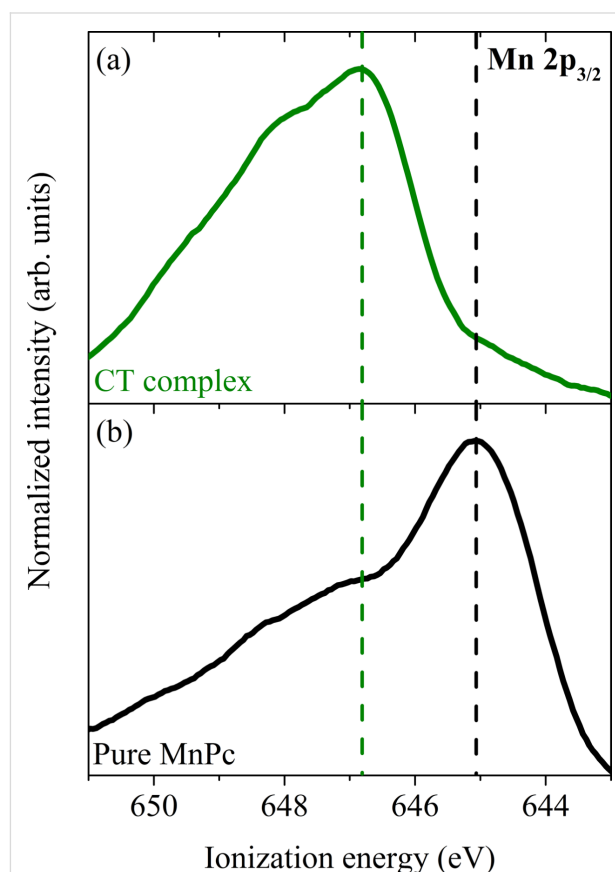


Figure 6: Core-level photoelectron spectroscopy data in the energy region of the Mn 2p_{3/2} core level (adapted from [104]). Note that the energy is referenced to the vacuum level in this case, since two different compounds with different Fermi-level positions are compared. Panel (a) shows the core-level emission of the MnPc/F₄TCNQ charge-transfer compound, while panel (b) shows that of pure MnPc. Clearly, the formation of MnPc/F₄TCNQ results in an energy shift of about 1.8 eV to higher ionization energy.

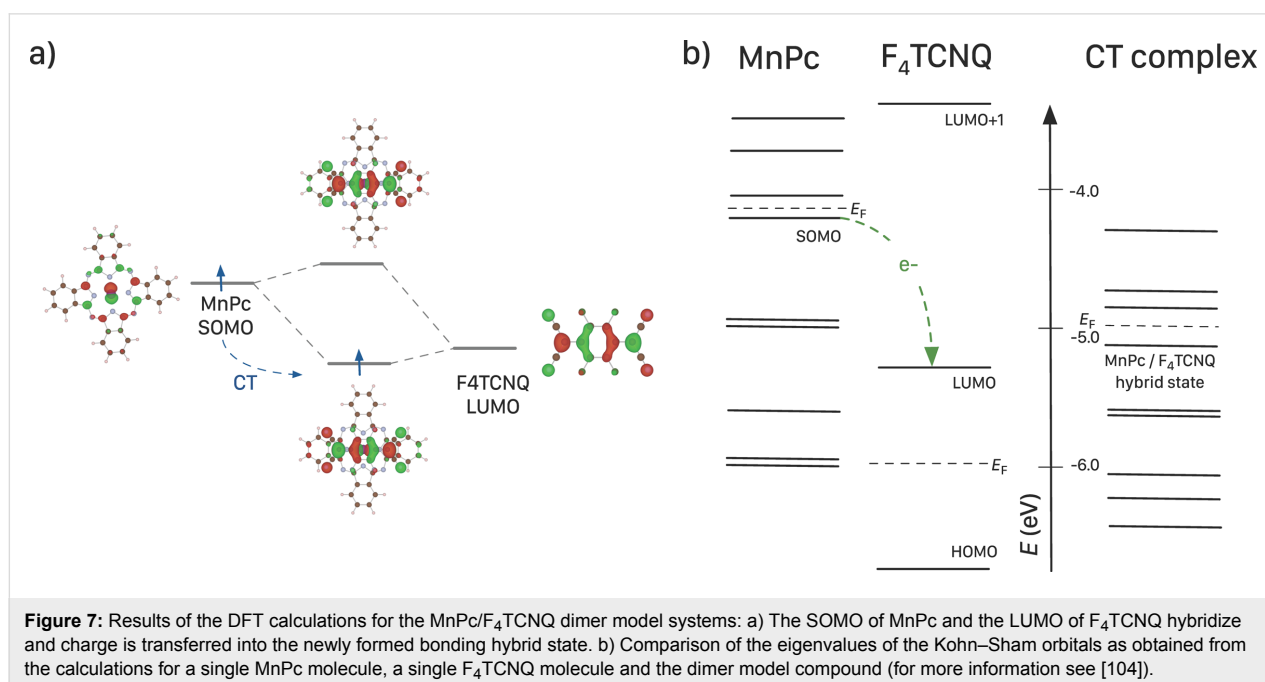
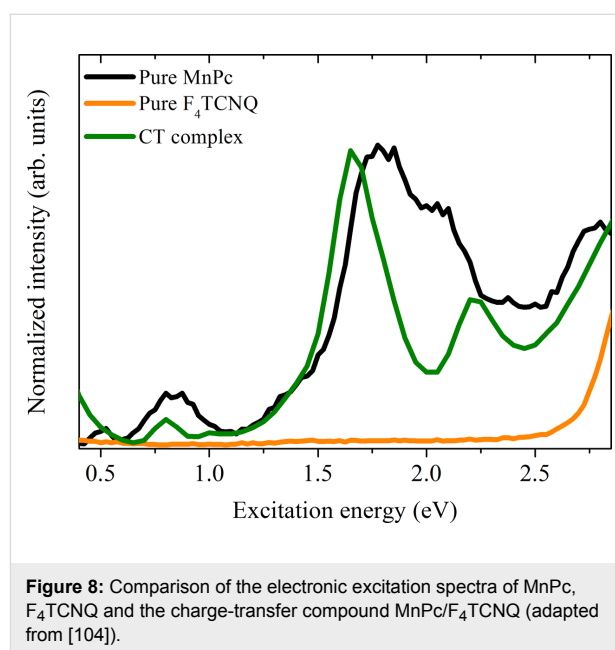
This figure presents the comparison of the Mn 2p_{3/2} core-level emission spectrum from a pure MnPc film and from the charge-transfer compound MnPc/F₄TCNQ. The overall spectral shape is controlled by an underlying multiplet structure [110,111] and is not discussed here further. For MnPc/F₄TCNQ the Mn 2p core level is significantly shifted to a higher ionization energy by about 1.8 eV, which results from the oxidation of the central Mn atom in MnPc. Shifts of the Mn 2p photoemission core-

level features to higher energies were also observed going from MnO to, e.g., Mn_3O_4 [112], where also the number of Mn 3d electrons is reduced. Thus, the core-level data corroborate our conclusion above.

The charge-transfer reaction between MnPc and F_4TCNQ was also studied on the basis of dimer calculations. These calculations indicate the formation of a hybrid state between the highest occupied orbital of MnPc (here called SOMO for singly occupied molecular orbital) and the LUMO of F_4TCNQ , see Figure 7. The energy position of this hybrid state is lower than that of the SOMO of MnPc. In addition, the calculations demonstrate a charge transfer between the molecules of about 0.6 electrons per dimer. The experimental results indicate an even larger charge transfer, which most likely is related to the localization error inherent to all DFT semi-local exchange correlation functionals [113]. Nevertheless, the calculations provide a reasonable understanding of the occupied electronic states of MnPc/ F_4TCNQ as demonstrated in Figure 5. Both the disappearance of the spectral feature at lowest binding energy of MnPc as well as the broadening of the structures is well reproduced. The calculations also indicate that the hybrid state is too close to the MnPc HOMO to be resolved spectroscopically.

The formation of the charge-transfer complex MnPc/ F_4TCNQ also results in corresponding changes of the electronic excitation spectra in comparison to those from pure MnPc or F_4TCNQ . This is illustrated in Figure 8, where the corresponding data from EELS are depicted. The measured spectrum of

pure F_4TCNQ is characterized by a rather large energy gap and an excitation onset at about 2.7 eV that is followed by a broad structure around 3.3 eV [114,115]. The rather complex excitation spectrum of MnPc was already discussed above. The formation of the charge-transfer compound MnPc/ F_4TCNQ gives rise to clearly different electronic excitations. The lowest excitation feature as seen for pure MnPc at about 0.5 eV cannot be seen any more. This can be associated to the removal of an electron from the leading orbital of MnPc as discussed above. The excitation spectrum of MnPc/ F_4TCNQ consists of rather sharp

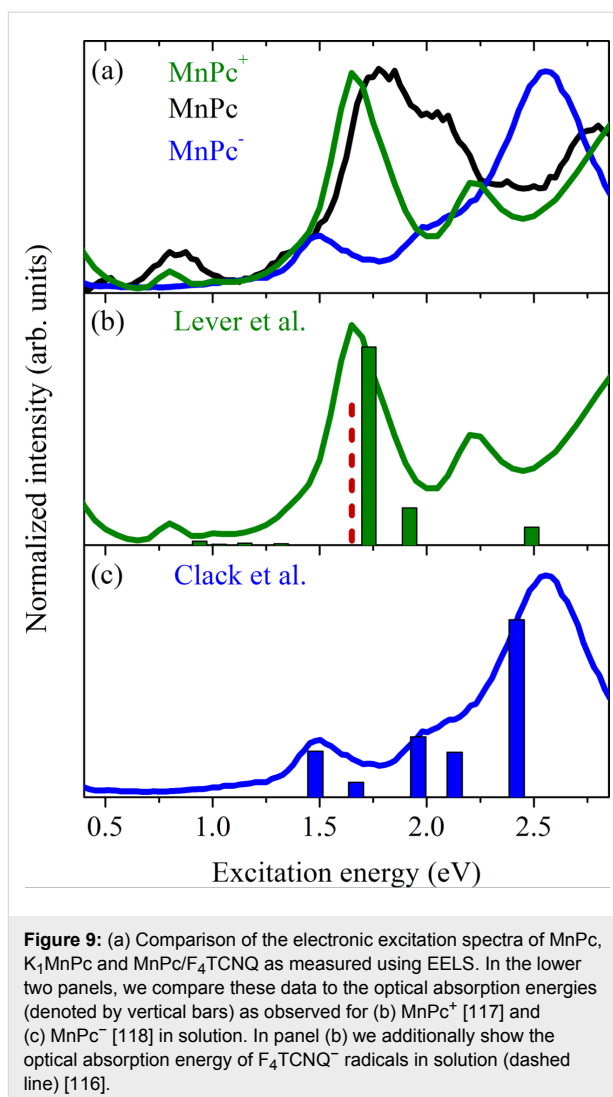


excitation maxima at around 0.8 eV, 1.65 eV and 2.2 eV. A detailed quantitative description of these excitations has not been achieved yet [104]. In general, our data show that the MnPc/F₄TCNQ compound has an energy gap of about 0.6 eV represented by the spectral onset in the excitation data. In particular, the lowest energy (gap) excitation is ascribed to the excitation within the two-level system, which originates from the charge-transfer reaction and the related hybrid-state formation as discussed above.

Thus, the oxidation of MnPc molecules upon the formation of the new charge transfer salt MnPc/F₄TCNQ is clearly seen in our spectroscopic data and supporting calculations. Together with the results on potassium-doped MnPc as presented in the previous chapter, this nicely demonstrates the variability of MnPc in charge-transfer compounds, where it can be either reduced or oxidized. We conclude the discussion of the MnPc-based charge-transfer compounds with a comparison of the electronic excitation spectra of MnPc, oxidized MnPc⁺ and reduced MnPc⁻ as measured for the MnPc/F₄TCNQ and K₁MnPc compounds. We argue that for both compounds the low-energy excitations predominantly stem from MnPc-derived orbitals. In the case of K₁MnPc this is quite clear since potassium ions do not contribute in the relevant energy region. For MnPc/F₄TCNQ the situation is more complex. Based on a purely ionic picture, also excitations from the negatively charged F₄TCNQ should show up. In solution, an excitation for F₄TCNQ⁻ radicals at about 1.65 eV was reported [116]. Keeping this in mind, we present a comparison of our electronic excitation data obtained using EELS for MnPc, K₁MnPc (MnPc⁻), and MnPc/F₄TCNQ (MnPc⁺) in Figure 9. In addition, we also included the energies of optical absorption studies for oxidized and reduced MnPc in solution [117,118]. Taking into account a broadening upon transition from single molecules in solution to the solid state as well as energy shifts due to different polarization screening, there is very good agreement between our data and those in solution from the literature. This nicely corroborates the interpretation of our results in terms of MnPc salts and the related electronic properties.

MnPc/F₆TCNNQ: charge transfer at an interface

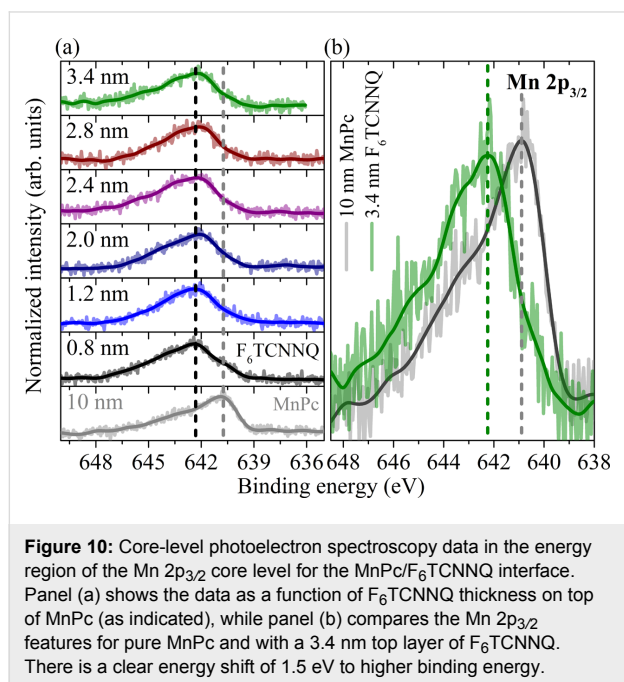
The organic heterojunction MnPc/F₆TCNNQ represents an example, where MnPc is involved in a charge transfer across an interface. In consideration of the results of the previous chapter, it is reasonable to expect charge transfer between MnPc and F₆TCNNQ, since F₆TCNNQ is an even stronger electron acceptor compared to F₄TCNQ. Moreover, it is larger and heavier, which prevents it from diffusion in or into organic films. Thus, it will form well-defined interfaces, i.e., also the charge-transfer reaction is confined to the interface region.



In general, charge transfer from insulators or semiconductors on one side of the interface to those on the other side can have dramatic effects and result in interfacial electronic properties that differ substantially from those of the individual materials [32,119-124]. For instance, particular interfaces between two initially semiconducting organic materials were shown to become even metallic [32,120].

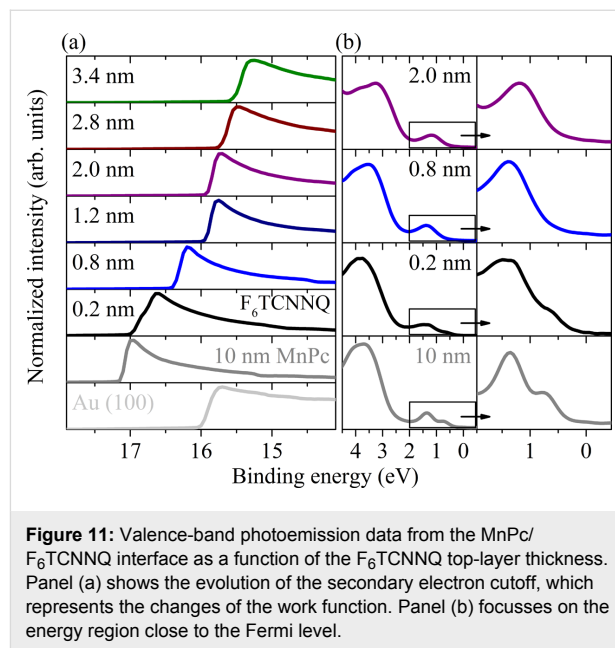
First photoemission studies of the MnPc/F₆TCNNQ interface indeed suggest that there is substantial charge transfer, which might also lead to interesting physics at these interfaces. In Figure 10 we show the results of the measurements of the Mn 2p_{3/2} core-level emission during the formation of the MnPc/F₆TCNNQ interface. Here, F₆TCNNQ was deposited stepwise onto a 10 nm thick MnPc film. Due to the rather small electron escape depth (a few angstroms only [39]) the Mn 2p_{3/2} core-level data increasingly stem from regions very close to the interface with increasing F₆TCNNQ top layer thickness. The

data in Figure 10 demonstrate a clear change in line shape and energy position of the the Mn $2p_{3/2}$ core-level feature, which is analogous to the changes seen above for MnPc/ F_4 TCNQ. Thus, these data evidence that there is charge transfer at the MnPc/ F_6 TCNNQ interface and that the Mn central atom in MnPc again is oxidized.



Furthermore, also the valence band data clearly indicate this charge transfer. Figure 11 illustrates the evolution of the high binding energy cutoff (a) and the energy region close to the Fermi energy (b). The data in panel (a) show the evolution of the work function of the layer system as a function of increasing F_6 TCNNQ layer thickness. With the exception of the thinnest F_6 TCNNQ layer on top of MnPc, the well-pronounced and sharp cutoff spectra affirm the formation of well-defined organic layers. In the case of 0.2 nm F_6 TCNNQ, a step is visible in the region of the secondary cutoff, which most likely is due to a coverage of less than a monolayer F_6 TCNNQ on MnPc, which results in surface/interface regions with and without the charge-transfer reaction. Figure 11b depicts the data close to the Fermi level for a selection of F_6 TCNNQ overlayer thicknesses. Again, for pure MnPc the spectrum is characterized by a two peak feature as described above. The features at lowest binding energy (about 0.7 eV) vanishes when F_6 TCNNQ is added, which signals the oxidation of MnPc in analogy to the previous section. The second feature, initially at about 1.4 eV, first broadens, then becomes somewhat sharper again and shifts to 1.2 eV. We attribute this feature to the emission from the now filled, formerly lowest unoccupied molecular orbital of F_6 TCNNQ, which is occupied at the interface as a result of the

charge transfer. Consequently, our data provide evidence for a considerable charge transfer at the MnPc/ F_6 TCNNQ interface, and further investigations are necessary to provide more insight into the physics as, e.g., whether there is electrical conduction along the interface.

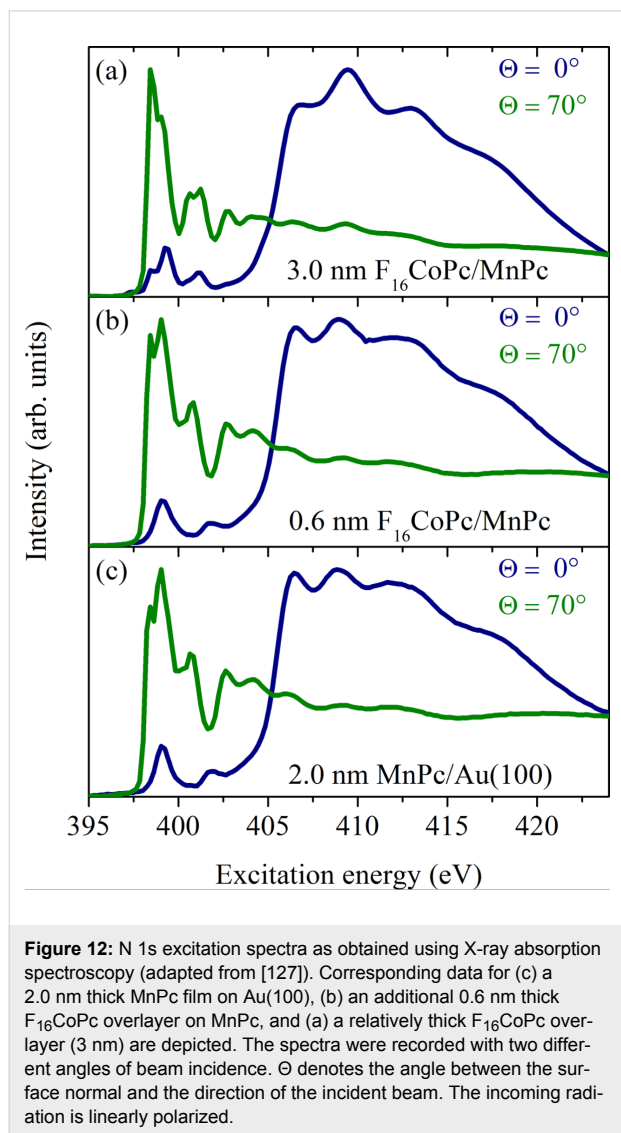


A charge and spin transfer interface: MnPc/ F_{16} CoPc

A further example in which a charge transfer across an interface results in new physical properties at this interface is provided by bringing together MnPc and F_{16} CoPc. Highly ordered interfaces of this kind were realized by depositing one of the two phthalocyanines on a gold(100) single crystal, which resulted in well-oriented thin films [125-127]. Subsequently, the partner phthalocyanine was deposited on top, which finally gave rise to well-ordered heterojunctions as revealed by polarization dependent X-ray absorption spectroscopy (XAS) studies [127].

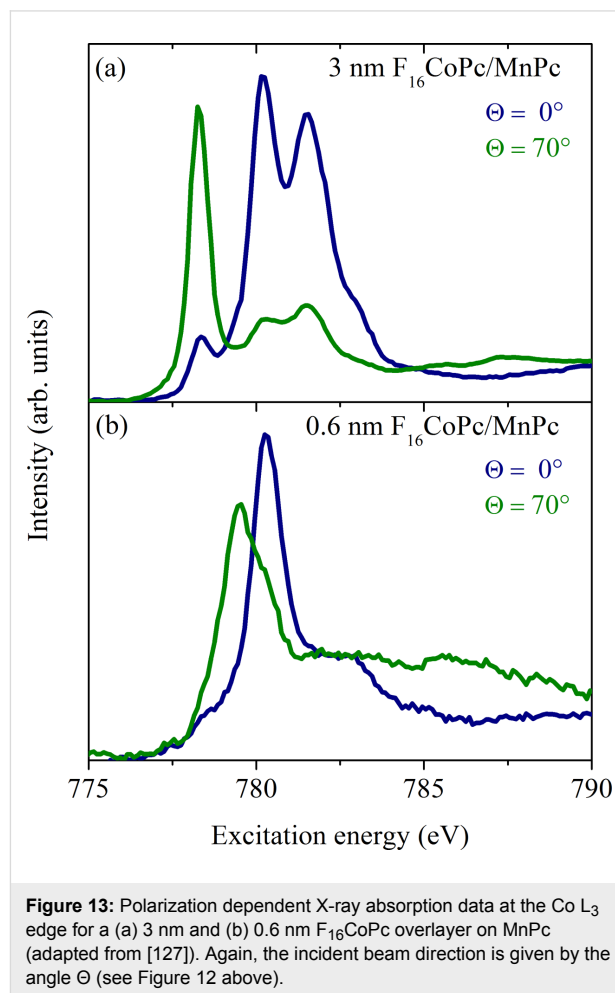
In Figure 12 we present corresponding N 1s absorption spectra for pure MnPc and two different film thicknesses of an F_{16} CoPc overlayer, deposited on a gold single crystal. Different light polarizations with respect to the film surface normal were achieved by variation of the angle between the incident light and the surface normal (see angles in Figure 12). For phthalocyanines it is well known that $1s \rightarrow \sigma^*$ and $1s \rightarrow \pi^*$ excitations take place for a light polarization vector perpendicular and parallel to the molecular planes, respectively. Moreover, previous studies [128,129] have demonstrated that the relatively sharp N 1s excitation features around 398 eV are due to transitions from the N 1s core level into the unoccupied π^* orbitals

with N 2p orbital contributions, which are oriented perpendicular to the molecular plane. The higher energy structures above 405 eV are related to N 1s→ σ^* transitions.



The data shown in Figure 12 reveal a very strong polarization dependence for the absorption edges of pure MnPc. The observed intensity variations show that the respective phthalocyanine molecules are arranged parallel to the substrate surface with a very high degree of orientation. The mean deviation from exactly parallel lying molecules is only about 5% [127]. This very high degree of order is also kept across the MnPc/F₁₆CoPc interface, as can be seen from Figure 12. In other words, the two phthalocyanines form an interface where they lie face to face. This high degree of order also allowed us to study the anisotropy of the excitations into Co-derived 3d levels in F₁₆CoPc close and far from the interface to MnPc. The corresponding data at the Co 2p absorption edges again show a very

clear anisotropy when taken with different polarization directions of the incident synchrotron radiation, as depicted in Figure 13.

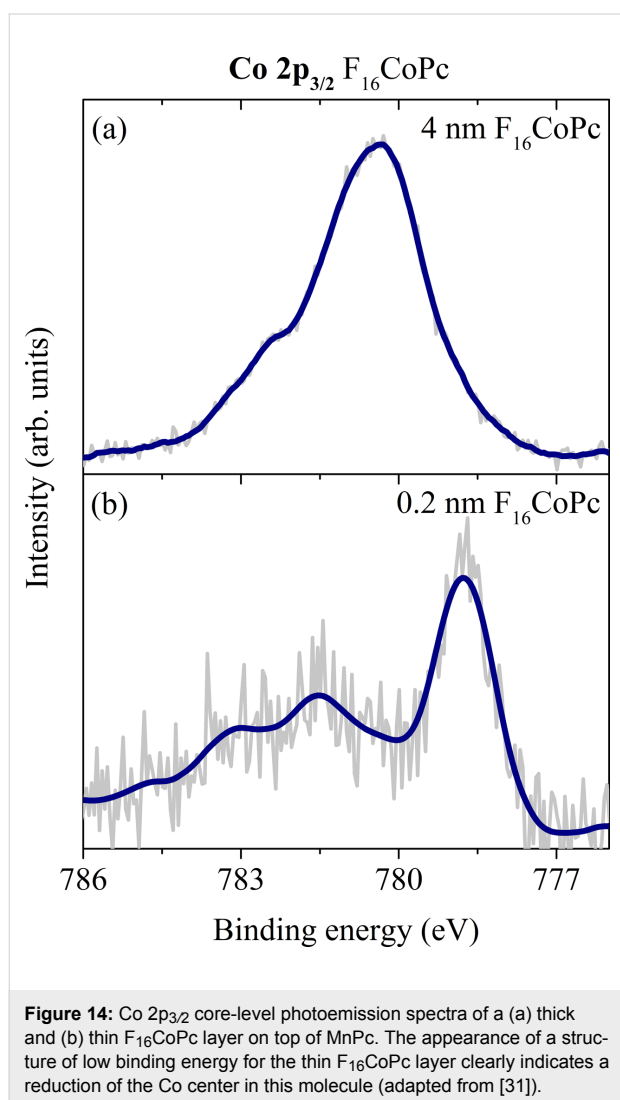


The data for a 3 nm thick F₁₆CoPc film on top of MnPc are very similar to the corresponding absorption spectra of pure CoPc on gold [130]. This indicates that fluorination of CoPc has little impact on the electronic 3d states of the central Co atom. The lowest lying absorption feature, which is maximal for a light polarization perpendicular to the F₁₆CoPc molecules, can be assigned to transitions from the Co 2p into unoccupied 3d_{z²} states. The higher lying features stem from a multiplet structure related to the excitations into the Co 3d_{x²-y²} orbital [130].

For thin F₁₆CoPc films deposited onto MnPc we observe considerably different spectra. Features that are characteristic for pure F₁₆CoPc disappear while new structures show up around 780 eV, which are still anisotropic. This provides clear evidence that there is a reaction between F₁₆CoPc and MnPc at the corresponding interface, which affects the cobalt states of F₁₆CoPc. This conclusion is supported by equivalent investiga-

tions of CoPc monolayers on gold and silver surfaces, where similar changes in the absorption spectra were observed and where a charge transfer between the cobalt 3d states and the underlying metallic substrate occurs [131,132]. Moreover, also the Mn $L_{2,3}$ absorption edge is subject to substantial changes for MnPc molecules in contact to F_{16} CoPc in comparison to pure MnPc [127]. This provides clear evidence that the charge transfer at the MnPc/ F_{16} CoPc interface again results in charge removal from Mn 3d orbitals in MnPc.

A charge-transfer reaction at the MnPc/ F_{16} CoPc interface is also seen in photoelectron spectroscopy studies [31]. Looking at the Co $2p_{3/2}$ core-level data of this interface, there is a significant change in line shape and binding energy as a function of the F_{16} CoPc layer thickness. These data are shown in Figure 14.



The data for the thin F_{16} CoPc layer are rather similar to what has been reported so far for cobalt porphyrines and phthalocyanines deposited on various metals. In these cases, a relatively strong interaction of the Co center of the molecules and the metal surface takes place [133–137]. The result of this interaction usually is rationalized in terms of a reduction of the metal center to Co(I). Consequently, there is clear evidence that at the MnPc/ F_{16} CoPc interface the Co center of F_{16} CoPc is also reduced as a result of a charge transfer across this interface. Also, a corresponding shift of the Mn 2p core-level feature to a higher binding energy is observed [31] (cf. Figure 6). The charge transfer arises from the formation of hybrid states between the transition-metal centers of the two phthalocyanines, with a concomitant oxidation of Mn-derived states of MnPc, similar to the charge transfer as discussed in previous sections. This situation is nicely supported by model calculations of a MnPc/ F_{16} CoPc dimer [31,138]. The calculations demonstrate that the states of the two phthalocyanines combine to form new bonding and anti-bonding states. The Mn $3d_{xz}$ and the Co $3d_{z^2}$ states hybridize and form a two-level system as illustrated qualitatively in Figure 15.

The occupation of the lower of these hybrid states also is related to the observed charge transfer. Intriguingly, as a result of our calculations the MnPc/ F_{16} CoPc dimer is characterized by a net spin of $S = 2$. Thus, the charge transfer is connected to a transfer/change of spin, which justifies to call the corresponding interface a spin-transfer interface with potential applications in the area of spintronics.

Finally, we note that also a bulk material consisting of MnPc/ F_{16} CoPc dimers could be prepared via the co-evaporation of these two materials. Spectroscopic studies of the resulting films confirmed the formation of MnPc/ F_{16} CoPc charge-transfer dimers in analogy to the related interface as discussed above [139]. The electronic excitation spectrum of these co-evaporated MnPc/ F_{16} CoPc films is characterized by a new feature at low energies (about 0.6 eV). Our density functional theory based calculations of the excitation spectrum reveal that this low-energy signal is due to transitions between the states of the dimer related two level system (see Figure 15).

Conclusion

The compilation of our results on bulk compounds and interfaces based on manganese phthalocyanines and partners, where the phthalocyanine is either reduced or oxidized demonstrates the variability of MnPc in the formation of novel, potentially interesting systems. Moreover, apart from interesting electronic properties that are associated with the charge transfer in either case, the spin/magnetic state of MnPc must also be changed since Mn 3d orbitals participate in the charge transfer. Future studies will certainly unravel more details and intriguing features in these respects.

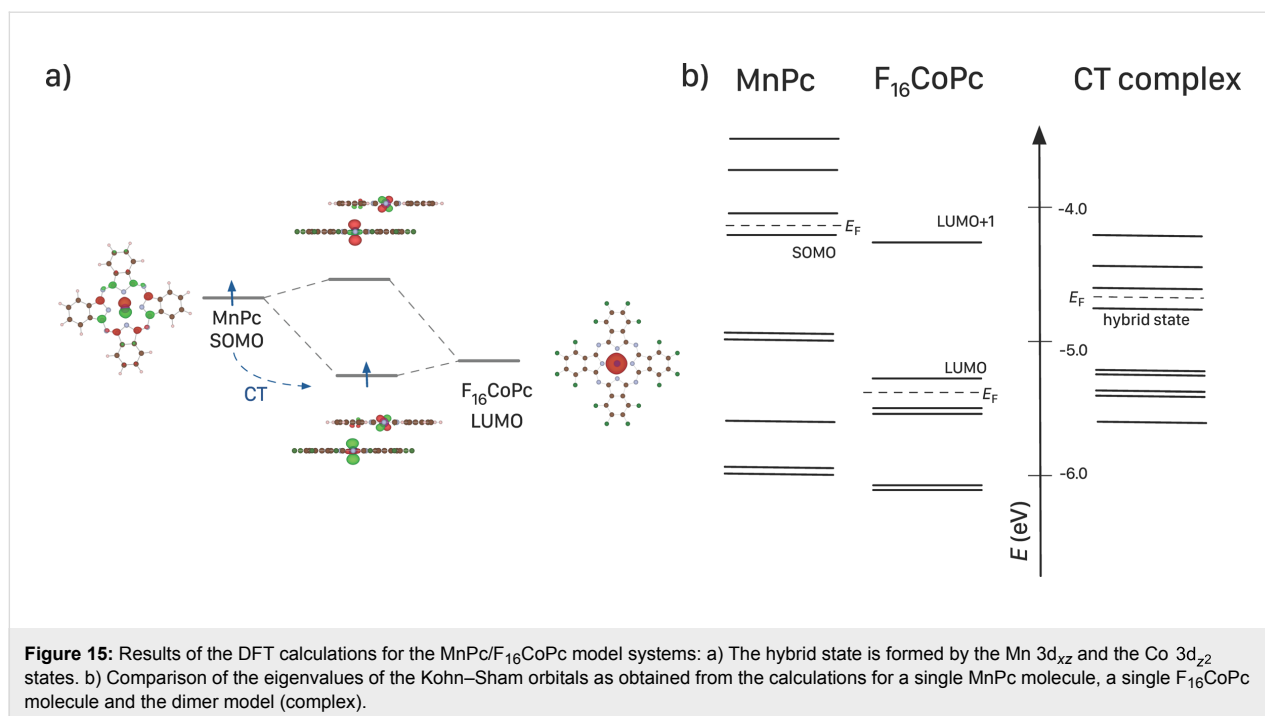


Figure 15: Results of the DFT calculations for the MnPc/F₁₆CoPc model systems: a) The hybrid state is formed by the Mn 3d_{xz} and the Co 3d_{z²} states. b) Comparison of the eigenvalues of the Kohn–Sham orbitals as obtained from the calculations for a single MnPc molecule, a single F₁₆CoPc molecule and the dimer model (complex).

Acknowledgements

We are grateful to R. Hübel, S. Leger and M. Naumann for technical assistance. Financial support by the Deutsche Forschungsgemeinschaft within the Forschergruppe FOR 1154 (KN393/14, KO1924/5, ZA146/23), as well as in projects KN393/25, and HA5070/3 is gratefully acknowledged.

References

- Gregory, P. J. *Porphyrins Phthalocyanines* **2000**, *4*, 432–437. doi:10.1002/(SICI)1099-1409(200006/07)4:4<432::AID-JPP254>3.0.CO;2-N
- Singh, T. B.; Sariciftci, N. S. *Annu. Rev. Mater. Res.* **2006**, *36*, 199–230. doi:10.1146/annurev.matsci.36.022805.094757
- Rand, B. P.; Genoe, J.; Heremans, P.; Poortmans, J. *Prog. Photovoltaics* **2007**, *15*, 659–676. doi:10.1002/pip.788
- Li, Y.; Chen, S.; Liu, Q.; Wang, L.; Someya, T.; Ma, J.; Wang, X.; Hu, Z. *J. Phys. Chem. C* **2012**, *116*, 4287–4292. doi:10.1021/jp210547j
- Lin, Y.; Li, Y.; Zhan, X. *Chem. Soc. Rev.* **2012**, *41*, 4245–4272. doi:10.1039/c2cs15313k
- Baeg, K.-J.; Binda, M.; Natali, D.; Caironi, M.; Noh, Y.-Y. *Adv. Mater.* **2013**, *25*, 4267–4295. doi:10.1002/adma.201204979
- Melville, O. A.; Lessard, B. H.; Bender, T. P. *ACS Appl. Mater. Interfaces* **2015**, *7*, 13105–13118. doi:10.1021/acsami.5b01718
- Gsänger, M.; Bialas, D.; Huang, L.; Stolte, M.; Würthner, F. *Adv. Mater.* **2016**, *28*, 3615–3645. doi:10.1002/adma.201505440
- van den Brink, J.; Morpurgo, A. F. *Nature* **2007**, *450*, 177–178. doi:10.1038/450177a
- Lach, S.; Altenhof, A.; Tarafder, K.; Schmitt, F.; Ali, M. E.; Vogel, M.; Sauther, J.; Oppeneer, P. M.; Ziegler, C. *Adv. Funct. Mater.* **2012**, *22*, 989–997. doi:10.1002/adfm.201102297
- Djegloul, F.; Ibrahim, F.; Cantoni, M.; Bowen, M.; Joly, L.; Boukari, S.; Ohresser, P.; Bertran, F.; Le Fèvre, P.; Thakur, P.; Scheurer, F.; Miyamachi, T.; Mattana, R.; Seneor, P.; Jaafar, A.; Rinaldi, C.; Javaid, S.; Arabski, J.; Kappler, J. P.; Wulfhekel, W.; Brookes, N. B.; Bertacco, R.; Taleb-Ibrahimi, A.; Alouani, M.; Beaurepaire, E.; Weber, W. *Sci. Rep.* **2013**, *3*, 1272. doi:10.1038/srep01272
- Barraud, C.; Bouzehouane, K.; Deranlot, C.; Kim, D. J.; Rakshit, R.; Shi, S.; Arabski, J.; Bowen, M.; Beaurepaire, E.; Boukari, S.; Petroff, F.; Seneor, P.; Mattana, R. *Dalton Trans.* **2016**, *45*, 16694–16699. doi:10.1039/C6DT02467J
- Barracough, C. G.; Martin, R. L.; Mitra, S.; Sherwood, R. C. *J. Chem. Phys.* **1970**, *53*, 1638–1642. doi:10.1063/1.1674236
- Mitra, S.; Gregson, A.; Hatfield, W. E.; Weller, R. R. *Inorg. Chem.* **1983**, *22*, 1729–1732. doi:10.1021/ic00154a007
- Heutz, S.; Mitra, C.; Wu, W.; Fisher, A.; Kerridge, A.; Stoneham, M.; Harker, A.; Gardener, J.; Tseng, H.-H.; Jones, T.; Renner, C.; Aeppli, G. *Adv. Mater.* **2007**, *19*, 3618–3622. doi:10.1002/adma.200701458
- Taguchi, Y.; Miyake, T.; Margadonna, S.; Kato, K.; Prassides, K.; Iwasa, Y. *J. Am. Chem. Soc.* **2006**, *128*, 3313–3323. doi:10.1021/ja0582657
- Mazur, U.; Hipps, K. J. *Phys. Chem. B* **1999**, *103*, 9721–9727. doi:10.1021/jp9923419
- Haidu, F.; Fechner, A.; Salvan, G.; Gordan, O. D.; Fronk, M.; Lehmann, D.; Mahns, B.; Knupfer, M.; Zahn, D. R. T. *AIP Adv.* **2013**, *3*, 062124. doi:10.1063/1.4812230
- Zahn, D. R. T.; Gavrilu, G. N.; Gorgoi, M. *Chem. Phys.* **2006**, *325*, 99–112. doi:10.1016/j.chemphys.2006.02.003
- Hill, I. G.; Kahn, A.; Soos, Z. G.; Pascal, R. A., Jr. *Chem. Phys. Lett.* **2000**, *327*, 181–188. doi:10.1016/S0009-2614(00)00882-4
- Brumboiu, I. E.; Totani, R.; de Simone, M.; Coreno, M.; Grazioli, C.; Lozzi, L.; Herper, H. C.; Sanyal, B.; Eriksson, O.; Puglia, C.; Brena, B. *J. Phys. Chem. A* **2014**, *118*, 927–932. doi:10.1021/jp4100747

22. Grobosch, M.; Mahns, B.; Loose, C.; Friedrich, R.; Schmidt, C.; Kortus, J.; Knupfer, M. *Chem. Phys. Lett.* **2011**, *505*, 122–125. doi:10.1016/j.cplett.2011.02.039
23. Kraus, R.; Grobosch, M.; Knupfer, M. *Chem. Phys. Lett.* **2009**, *469*, 121–124. doi:10.1016/j.cplett.2008.12.090
24. Rangger, G. M.; Hofmann, O. T.; Romaner, L.; Heimel, G.; Bröker, B.; Blum, R.-P.; Johnson, R. L.; Koch, N.; Zojer, E. *Phys. Rev. B* **2009**, *79*, 165306. doi:10.1103/PhysRevB.79.165306
25. Toyota, N.; Müller, J.; Lang, M. *Low-dimensional molecular metals*; Springer Science & Business Media, 2007. doi:10.1007/978-3-540-49576-5
26. Ouahab, L.; Yagubskii, E. *Organic conductors, superconductors and magnets: from synthesis to molecular electronics*; Springer Science & Business Media, 2004. doi:10.1007/978-94-007-1027-6
27. Walzer, K.; Maennig, B.; Pfeiffer, M.; Leo, K. *Chem. Rev.* **2007**, *107*, 1233–1271. doi:10.1021/cr050156n
28. Aziz, E. F.; Vollmer, A.; Eisebitt, S.; Eberhardt, W.; Pingel, P.; Neher, D.; Koch, N. *Adv. Mater.* **2007**, *19*, 3257–3260. doi:10.1002/adma.200700926
29. Lüssem, B.; Riede, M.; Leo, K. *Phys. Status Solidi A* **2013**, *210*, 9–43. doi:10.1002/pssa.201228310
30. Lüssem, B.; Tietze, M. L.; Kleemann, H.; Hoßbach, C.; Bartha, J. W.; Zakhidov, A.; Leo, K. *Nat. Commun.* **2013**, *4*, 2775. doi:10.1038/ncomms3775
31. Lindner, S.; Knupfer, M.; Friedrich, R.; Hahn, T.; Kortus, J. *Phys. Rev. Lett.* **2012**, *109*, 027601. doi:10.1103/PhysRevLett.109.027601
32. Krupskaya, Y.; Rückerl, F.; Knupfer, M.; Morpurgo, A. F. *Adv. Mater. Interfaces* **2016**, *3*, 1500863. doi:10.1002/admi.201500863
33. Rückerl, F.; Waas, D.; Büchner, B.; Knupfer, M. *J. Electron Spectrosc. Relat. Phenom.* **2017**, *215*, 1–7. doi:10.1016/j.elspec.2016.11.013
34. Toader, M.; Gopakumar, T. G.; Shukryna, P.; Hietschold, M. *J. Phys. Chem. C* **2010**, *114*, 21548–21554. doi:10.1021/jp1078295
35. Pinto, H.; Jones, R.; Goss, J. P.; Briddon, P. R. *J. Phys.: Condens. Matter* **2009**, *21*, 402001. doi:10.1088/0953-8984/21/40/402001
36. Méndez, H.; Heimel, G.; Winkler, S.; Frisch, J.; Opitz, A.; Sauer, K.; Wegner, B.; Oehzelt, M.; Röthel, C.; Duhm, S.; Többsens, D.; Koch, N.; Salzmann, I. *Nat. Commun.* **2015**, *6*, 8560. doi:10.1038/ncomms9560
37. Cardona, M.; Ley, L. *Photoemission in Solids I: General Principles*; Springer, 1978. doi:10.1007/3-540-08685-4
38. Ley, L.; Cardona, M. *Photoemission in Solids II: Case Studies*; Springer, 1979. doi:10.1007/3-540-09202-1
39. Hüfner, S. *Photoelectron spectroscopy: principles and applications*; Springer Science & Business Media, 2013.
40. Smith, N. V. *Rep. Prog. Phys.* **1988**, *51*, 1227. doi:10.1088/0034-4885/51/9/003
41. Himpfel, F.; Fauster, T. *J. Vac. Sci. Technol., A* **1984**, *2*, 815–821. doi:10.1116/1.572514
42. Fuggle, J. C.; Inglesfield, J. E. Introduction. *Unoccupied Electronic States*; Springer, 1992; pp 1–23. doi:10.1007/3540541624_11
43. Fink, J. *Adv. Electron. Electron Phys.* **1989**, *75*, 121–232. doi:10.1016/S0065-2539(08)60947-6
44. Fink, J. Transmission electron energy-loss spectroscopy. *Unoccupied Electronic States*; Springer, 1992; pp 203–241. doi:10.1007/3540541624_17
45. Roth, F.; König, A.; Fink, J.; Büchner, B.; Knupfer, M. *J. Electron Spectrosc. Relat. Phenom.* **2014**, *195*, 85–95. doi:10.1016/j.elspec.2014.05.007
46. Knupfer, M.; Pichler, T.; Golden, M. S.; Fink, J.; Murgia, M.; Michel, R. H.; Zamboni, R.; Taliani, C. *Phys. Rev. Lett.* **1999**, *83*, 1443–1446. doi:10.1103/PhysRevLett.83.1443
47. Tompkins, H. G.; Irene, E. A. *Handbook of Ellipsometry*; William Andrew: New York, 2005. doi:10.1016/B978-081551499-2.50002-2
48. Fujiwara, H. *Spectroscopic Ellipsometry*; Wiley: Chichester, 2007. doi:10.1002/9780470060193
49. Haidu, F. Tailoring the Electronic and Optical Properties of Molecular Thin Films by Reducing and Oxidising Agents. Ph.D. Thesis, Technische Universität Chemnitz, Germany, 2014.
50. Stöhr, J. *NEXAFS Spectroscopy*; Springer, 1992. doi:10.1007/978-3-662-02853-7
51. Pederson, M. R.; Jackson, K. A. *Phys. Rev. B* **1990**, *41*, 7453–7461. doi:10.1103/PhysRevB.41.7453
52. Pederson, M. R.; Porezag, D. V.; Kortus, J.; Patton, D. C. *Phys. Status Solidi B* **2000**, *217*, 197–218. doi:10.1002/(SICI)1521-3951(200001)217:1<197::AID-PSSB197>3.0.CO;2-B
53. Porezag, D.; Pederson, M. R. *Phys. Rev. A* **1999**, *60*, 2840–2847. doi:10.1103/PhysRevA.60.2840
54. Perdew, J. P.; Burke, K.; Ernzerhof, M. *Phys. Rev. Lett.* **1996**, *77*, 3865–3868. doi:10.1103/PhysRevLett.77.3865
55. Grimme, S.; Antony, J.; Ehrlich, S.; Krieg, H. *J. Chem. Phys.* **2010**, *132*, 154104. doi:10.1063/1.3382344
56. Mahns, B.; Roth, F.; Grobosch, M.; Zahn, D. R. T.; Knupfer, M. *J. Chem. Phys.* **2011**, *134*, 194504. doi:10.1063/1.3591347
57. Eastwood, D.; Edwards, L.; Gouterman, M.; Steinfeld, J. *J. Mol. Spectrosc.* **1966**, *20*, 381–390. doi:10.1016/0022-2852(66)90009-9
58. Leznoff, C. C.; Lever, A. B. P., Eds. *Phthalocyanines: Properties and Applications*; VCH Publishers: New York, NY, USA, 1993.
59. Dini, D.; Hanack, M. *J. Porphyrins Phthalocyanines* **2004**, *8*, 915–933. doi:10.1142/S1088424604000301
60. Lucia, E. A.; Verderame, F. D. *J. Chem. Phys.* **1968**, *48*, 2674–2681. doi:10.1063/1.1669501
61. Hernandez, J. P.; Choi, S.-i. *J. Chem. Phys.* **1969**, *50*, 1524–1532. doi:10.1063/1.1671237
62. Hollebone, B. R.; Stillman, M. J. *Chem. Phys. Lett.* **1974**, *29*, 284–286. doi:10.1016/0009-2614(74)85032-3
63. Saito, T.; Sisk, W.; Kobayashi, T.; Suzuki, S.; Iwayanagi, T. *J. Phys. Chem.* **1993**, *97*, 8026–8031. doi:10.1021/j100132a036
64. Umeda, M.; Mohamedi, M.; Itoh, T.; Uchida, I. *J. Appl. Phys.* **2001**, *90*, 3984–3987. doi:10.1063/1.1403674
65. Auerhammer, J. M.; Knupfer, M.; Peisert, H.; Fink, J. *Surf. Sci.* **2002**, *506*, 333–338. doi:10.1016/S0039-6028(02)01517-0
66. Knupfer, M.; Schwieger, T.; Peisert, H.; Fink, J. *Phys. Rev. B* **2004**, *69*, 165210. doi:10.1103/PhysRevB.69.165210
67. Maslov, V. G. *Opt. Spectrosc.* **2006**, *101*, 853–861. doi:10.1134/S0030400X0612006X
68. Bondarev, I. V.; Popescu, A.; Younts, R. A.; Hoffman, B.; McAfee, T.; Dougherty, D. B.; Gundogdu, K.; Ade, H. W. *Appl. Phys. Lett.* **2016**, *109*, 213302. doi:10.1063/1.4968821
69. Fielding, P. E.; MacKay, A. G. *Aust. J. Chem.* **1975**, *28*, 1445–1454. doi:10.1071/CH9751445
70. Williamson, B. E.; VanCott, T. C.; Boyle, M. E.; Misener, G. C.; Stillman, M. J.; Schatz, P. N. *J. Am. Chem. Soc.* **1992**, *114*, 2412–2419. doi:10.1021/ja00033a016

71. Nyokong, T. *Struct. Bonding* **2010**, *135*, 45–87. doi:10.1007/978-3-642-04752-7_2
72. Stradi, D.; Diaz, C.; Martin, F.; Alcamí, M. *Theor. Chem. Acc.* **2011**, *128*, 497–503. doi:10.1007/s00214-010-0852-1
73. Friedrich, R.; Hahn, T.; Kortus, J.; Fronk, M.; Haidu, F.; Salvan, G.; Zahn, D. R. T.; Schlesinger, M.; Mehring, M.; Roth, F.; Mahns, B.; Knupfer, M. *J. Chem. Phys.* **2012**, *136*, 064704. doi:10.1063/1.3683253
74. Giovanelli, L.; Vilmercati, P.; Castellarin-Cudia, C.; Themlin, J.-M.; Porte, L.; Goldoni, A. *J. Chem. Phys.* **2007**, *126*, 044709. doi:10.1063/1.2432115
75. Flatz, K.; Grobosch, M.; Knupfer, M. *J. Chem. Phys.* **2007**, *126*, 214702. doi:10.1063/1.2741539
76. Roth, F.; König, A.; Kraus, R.; Knupfer, M. *J. Chem. Phys.* **2008**, *128*, 194711. doi:10.1063/1.2920179
77. de Groot, F. M. F.; Fuggle, J. C.; Thole, B. T.; Sawatzky, G. A. *Phys. Rev. B* **1990**, *41*, 928–937. doi:10.1103/PhysRevB.41.928
78. Molodtsova, O. V.; Knupfer, M.; Aristov, V. Yu.; Vyalikh, D. V.; Zhilin, V. M.; Ossipyan, Yu. A. *J. Appl. Phys.* **2008**, *103*, 053711. doi:10.1063/1.2874001
79. Aristov, V. Yu.; Molodtsova, O. V.; Maslyuk, V. V.; Vyalikh, D. V.; Bredow, T.; Mertig, I.; Preobrajenski, A. B.; Knupfer, M. *Org. Electron.* **2010**, *11*, 1461–1468. doi:10.1016/j.orgel.2010.04.028
80. Nilson, K.; Åhlund, J.; Shariati, M.-N.; Schiessling, J.; Palmgren, P.; Brena, B.; Göthelid, E.; Hennies, F.; Huismans, Y.; Evangelista, F.; Rudolf, P.; Göthelid, M.; Mårtensson, N.; Puglia, C. *J. Chem. Phys.* **2012**, *137*, 044708. doi:10.1063/1.4738755
81. Peisert, H.; Knupfer, M.; Fink, J. *Surf. Sci.* **2002**, *515*, 491–498. doi:10.1016/S0039-6028(02)01967-2
82. Grobosch, M.; Schmidt, C.; Kraus, R.; Knupfer, M. *Org. Electron.* **2010**, *11*, 1483–1488. doi:10.1016/j.orgel.2010.06.006
83. Aristov, V. Yu.; Molodtsova, O. V.; Knupfer, M. *Org. Electron.* **2011**, *12*, 372–375. doi:10.1016/j.orgel.2010.12.003
84. Marom, N.; Kronik, L. *Appl. Phys. A* **2009**, *95*, 165–172. doi:10.1007/s00339-008-5005-1
85. Giovannetti, G.; Brocks, G.; van den Brink, J. *Phys. Rev. B* **2008**, *77*, 035133. doi:10.1103/PhysRevB.77.035133
86. Wu, X.; Xu, C.; Wang, K.; Xiao, X. *J. Phys. Chem. C* **2016**, *120*, 15446–15452. doi:10.1021/acs.jpcc.6b03686
87. Phan, Q. T. N.; Heguri, S.; Tamura, H.; Nakano, T.; Nozue, Y.; Tanigaki, K. *Phys. Rev. B* **2016**, *93*, 075130. doi:10.1103/PhysRevB.93.075130
88. Ruff, A.; Sing, M.; Claessen, R.; Lee, H.; Tomić, M.; Jeschke, H. O.; Valentí, R. *Phys. Rev. Lett.* **2013**, *110*, 216403. doi:10.1103/PhysRevLett.110.216403
89. Caputo, M.; Di Santo, G.; Parisse, P.; Petaccia, L.; Floreano, L.; Verdini, A.; Panighel, M.; Struzzi, C.; Taleatu, B.; Lal, C.; Goldoni, A. *J. Phys. Chem. C* **2012**, *116*, 19902–19908. doi:10.1021/jp306640z
90. Mahns, B.; Roth, F.; Knupfer, M. *J. Chem. Phys.* **2012**, *136*, 134503. doi:10.1063/1.3699188
91. Knupfer, M.; Peisert, H.; Schwiager, T. *Phys. Rev. B* **2001**, *65*, 033204. doi:10.1103/PhysRevB.65.033204
92. Brühwiler, P. A.; Maxwell, A. J.; Nilsson, A.; Mårtensson, N.; Gunnarsson, O. *Phys. Rev. B* **1993**, *48*, 18296–18299. doi:10.1103/PhysRevB.48.18296
93. Knupfer, M.; Poirier, D. M.; Weaver, J. H. *Phys. Rev. B* **1994**, *49*, 2281–2284. doi:10.1103/PhysRevB.49.2281
94. Lof, R. W.; van Veenendaal, M. A.; Koopmans, B.; Jonkman, H. T.; Sawatzky, G. A. *Phys. Rev. Lett.* **1992**, *68*, 3924–3927. doi:10.1103/PhysRevLett.68.3924
95. Hotta, S.; Kobayashi, H. *Synth. Met.* **1994**, *66*, 117–122. doi:10.1016/0379-6779(94)90087-6
96. Sato, A.; Okada, M.; Saito, K.; Sorai, M. *Acta Crystallogr., Sect. C* **2001**, *57*, 564–565. doi:10.1107/S0108270100019454
97. Zhu, L.; Yi, Y.; Li, Y.; Kim, E.-G.; Coropceanu, V.; Brédas, J.-L. *J. Am. Chem. Soc.* **2012**, *134*, 2340–2347. doi:10.1021/ja210284s
98. Mahns, B.; Kataeva, O.; Islamov, D.; Hampel, S.; Steckel, F.; Hess, C.; Knupfer, M.; Büchner, B.; Himcinschi, C.; Hahn, T.; Renger, R.; Kortus, J. *Cryst. Growth Des.* **2014**, *14*, 1338–1346. doi:10.1021/cg401841n
99. Yoshida, Y.; Kumagai, Y.; Mizuno, M.; Isomura, K.; Nakamura, Y.; Kishida, H.; Saito, G. *Cryst. Growth Des.* **2015**, *15*, 5513–5518. doi:10.1021/acs.cgd.5b01138
100. Kataeva, O.; Khrižanforov, M.; Budnikova, Y.; Islamov, D.; Burganov, T.; Vandyukov, A.; Lyssenko, K.; Mahns, B.; Nohr, M.; Hampel, S.; Knupfer, M. *Cryst. Growth Des.* **2015**, *16*, 331–338. doi:10.1021/acs.cgd.5b01301
101. Chernenkaya, A.; Morherr, A.; Backes, S.; Popp, W.; Witt, S.; Kozina, X.; Nepijko, S. A.; Bolte, M.; Medjanik, K.; Öhrwall, G.; Krellner, C.; Baumgarten, M.; Elmers, H. J.; Schönhense, G.; Jeschke, H. O.; Valent, R. *J. Chem. Phys.* **2016**, *145*, 034702. doi:10.1063/1.4958659
102. Morherr, A.; Witt, S.; Chernenkaya, A.; Bäcker, J.-P.; Schönhense, G.; Bolte, M.; Krellner, C. *Phys. B (Amsterdam, Neth.)* **2016**, *496*, 98–105. doi:10.1016/j.physb.2016.05.023
103. Hu, P.; Du, K.; Wei, F.; Jiang, H.; Kloc, C. *Cryst. Growth Des.* **2016**, *16*, 3019–3027. doi:10.1021/acs.cgd.5b01675
104. Ruckerl, F.; Mahns, B.; Dodbiba, E.; Nikolis, V.; Herzig, M.; Büchner, B.; Knupfer, M.; Hahn, T.; Kortus, J. *J. Chem. Phys.* **2016**, *145*, 114702. doi:10.1063/1.4962578
105. Poirier, D. M. *Appl. Phys. Lett.* **1994**, *64*, 1356–1358. doi:10.1063/1.111933
106. Knupfer, M.; Poirier, D. M.; Weaver, J. H. *Phys. Rev. B* **1994**, *49*, 8464–8474. doi:10.1103/PhysRevB.49.8464
107. Poirier, D. M.; Olson, C. G.; Weaver, J. H. *Phys. Rev. B* **1995**, *52*, R11662–R11664. doi:10.1103/PhysRevB.52.R11662
108. Poirier, D. M.; Ownes, D. W.; Weaver, J. H. *Phys. Rev. B* **1995**, *51*, 1830–1843. doi:10.1103/PhysRevB.51.1830
109. Andjelković, L.; Stepanović, S.; Vlahović, F.; Zlatar, M.; Gruden, M. *Phys. Chem. Chem. Phys.* **2016**, *18*, 29122–29130. doi:10.1039/C6CP03859J
110. Nesbitt, H. W.; Banerjee, D. *Am. Mineral.* **1998**, *83*, 305–315. doi:10.2138/am-1998-3-414
111. Kröll, T.; Kraus, R.; Schönfelder, R.; Aristov, V. Yu.; Molodtsova, O.; Hoffmann, P.; Knupfer, M. *J. Chem. Phys.* **2012**, *137*, 054306. doi:10.1063/1.4738754
112. Oku, M.; Hirokawa, K.; Ikeda, S. *J. Electron Spectrosc. Relat. Phenom.* **1975**, *7*, 465–473. doi:10.1016/0368-2048(75)85010-9
113. Mori-Sánchez, P.; Cohen, A. J.; Yang, W. *Phys. Rev. Lett.* **2008**, *100*, 146401. doi:10.1103/PhysRevLett.100.146401
114. Dixon, D. A.; Calabrese, J. C.; Miller, J. S. *J. Phys. Chem.* **1989**, *93*, 2284–2291. doi:10.1021/j100343a019
115. Müller, E.; Mahns, B.; Büchner, B.; Knupfer, M. *J. Chem. Phys.* **2015**, *142*, 184702. doi:10.1063/1.4919881
116. Le, T. H.; Lu, J.; Bond, A. M.; Martin, L. L. *Inorg. Chim. Acta* **2013**, *395*, 252–254. doi:10.1016/j.ica.2012.10.019
117. Lever, A. B. P.; Pickens, S. R.; Minor, P. C.; Licoccia, S.; Ramaswamy, B. S.; Magnell, K. *J. Am. Chem. Soc.* **1981**, *103*, 6800–6806. doi:10.1021/ja00413a003

118. Clack, D. W.; Yandle, J. R. *Inorg. Chem.* **1972**, *11*, 1738–1742. doi:10.1021/ic50114a003
119. Wang, J.; Wang, H.; Yan, X.; Huang, H.; Yan, D. *Appl. Phys. Lett.* **2005**, *87*, 093507. doi:10.1063/1.2037204
120. Alves, H.; Molinari, A. S.; Xie, H.; Morpurgo, A. F. *Nat. Mater.* **2008**, *7*, 574–580. doi:10.1038/nmat2205
121. Nakano, M.; Alves, H.; Molinari, A. S.; Ono, S.; Minder, N.; Morpurgo, A. F. *Appl. Phys. Lett.* **2010**, *96*, 232102. doi:10.1063/1.3449558
122. Lezama, I. G.; Nakano, M.; Minder, N. A.; Chen, Z.; Di Girolamo, F. V.; Facchetti, A.; Morpurgo, A. F. *Nat. Mater.* **2012**, *11*, 788–794. doi:10.1038/nmat3383
123. Alves, H.; Pinto, R. M.; Maçôas, E. S. *Nat. Commun.* **2013**, *4*, 1842. doi:10.1038/ncomms2890
124. Krupskaya, Y.; Lezama, I. G.; Morpurgo, A. F. *Adv. Funct. Mater.* **2016**, *26*, 2334–2340. doi:10.1002/adfm.201502082
125. Molodtsova, O. V.; Knupfer, M.; Ossipyan, Yu. A.; Aristov, V. Yu. *J. Appl. Phys.* **2008**, *104*, 083704. doi:10.1063/1.3000105
126. Petraki, F.; Peisert, H.; Hoffmann, P.; Uihlein, J.; Knupfer, M.; Chassé, T. *J. Phys. Chem. C* **2012**, *116*, 5121–5127. doi:10.1021/jp211445n
127. Lindner, S.; Mahns, B.; Treske, U.; Vilkov, O.; Haidu, F.; Fronk, M.; Zahn, D. R. T.; Knupfer, M. *J. Chem. Phys.* **2014**, *141*, 094706. doi:10.1063/1.4894757
128. Ahlund, J.; Nilson, K.; Schiessling, J.; Kjeldgaard, L.; Berner, S.; Mårtensson, N.; Puglia, C.; Brena, B.; Nyberg, M.; Luo, Y. *J. Chem. Phys.* **2006**, *125*, 034709. doi:10.1063/1.2212404
129. Shariati, M.-N.; Lüder, J.; Bidermane, I.; Ahmadi, S.; Göthelid, E.; Palmgren, P.; Sanyal, B.; Eriksson, O.; Piancastelli, M. N.; Brena, B.; Puglia, C. *J. Phys. Chem. C* **2013**, *117*, 7018–7025. doi:10.1021/jp307626n
130. Kroll, T.; Aristov, V. Yu.; Molodtsova, O. V.; Ossipyan, Yu. A.; Vyalikh, D. V.; Büchner, B.; Knupfer, M. *J. Phys. Chem. A* **2009**, *113*, 8917–8922. doi:10.1021/jp903001v
131. Petraki, F.; Peisert, H.; Biswas, I.; Aygül, U.; Latteyer, F.; Vollmer, A.; Chassé, T. *J. Phys. Chem. Lett.* **2010**, *1*, 3380–3384. doi:10.1021/jz101395s
132. Petraki, F.; Peisert, H.; Latteyer, F.; Aygül, U.; Vollmer, A.; Chassé, T. *J. Phys. Chem. C* **2011**, *115*, 21334–21340. doi:10.1021/jp207568q
133. Bai, Y.; Buchner, F.; Kellner, I.; Schmid, M.; Vollnhals, F.; Steinrück, H.-P.; Marbach, H.; Gottfried, J. M. *New J. Phys.* **2009**, *11*, 125004. doi:10.1088/1367-2630/11/12/125004
134. Petraki, F.; Peisert, H.; Biswas, I.; Chassé, T. *J. Phys. Chem. C* **2010**, *114*, 17638–17643. doi:10.1021/jp104141s
135. Lindner, S.; Treske, U.; Grobosch, M.; Knupfer, M. *Appl. Phys. A* **2011**, *105*, 921–925. doi:10.1007/s00339-011-6648-x
136. Lindner, S.; Treske, U.; Knupfer, M. *Appl. Surf. Sci.* **2013**, *267*, 62–65. doi:10.1016/j.apsusc.2012.06.104
137. Gottfried, J. M. *Surf. Sci. Rep.* **2015**, *70*, 259–379. doi:10.1016/j.surfrep.2015.04.001
138. Friedrich, R.; Lindner, S.; Hahn, T.; Loose, C.; Liebing, S.; Knupfer, M.; Kortus, J. *Phys. Rev. B* **2013**, *87*, 115423. doi:10.1103/PhysRevB.87.115423
139. Lindner, S.; Mahns, B.; König, A.; Roth, F.; Knupfer, M.; Friedrich, R.; Hahn, T.; Kortus, J. *J. Chem. Phys.* **2013**, *138*, 024707. doi:10.1063/1.4774060

License and Terms

This is an Open Access article under the terms of the Creative Commons Attribution License (<http://creativecommons.org/licenses/by/4.0>), which permits unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

The license is subject to the *Beilstein Journal of Nanotechnology* terms and conditions: (<http://www.beilstein-journals.org/bjnano>)

The definitive version of this article is the electronic one which can be found at: [doi:10.3762/bjnano.8.160](https://doi.org/10.3762/bjnano.8.160)